

20.1 The experimental gas laws Page 319

4. a)

BEFORE

AFTER

$V_h = ?$	$P_i = 110 \text{ kPa}$
$V_h = \text{Volume of head space}$	$V_f = 1.20 \times 10^{-4} \text{ m}^3$
$V_t = \text{Volume of flask}$	$V_t = V_h + V_f$

$P_f = 135 \text{ kPa}$
$V_f = 1.20 \times 10^{-4}$
$V_t = V_f$

T is constant

$PV = \text{constant}$

$$\begin{aligned}
 P_i V_t &= P_f V_t \\
 110(V_h + 1.20 \times 10^{-4}) &= 135(1.20 \times 10^{-4}) \\
 110 V_h + 110(1.20 \times 10^{-4}) &= 135(1.20 \times 10^{-4}) \\
 110 V_h + 0.0132 &= 0.0162 \\
 110 V_h &= 3 \times 10^{-3} \\
 V_h &= \frac{3 \times 10^{-3}}{110} = 27.27 \times 10^{-6} \text{ m}^3 \\
 &= 2.7 \times 10^{-5} \text{ m}^3 \text{ to 2 SF}
 \end{aligned}$$

b)

V_h	$P_i = 110 \text{ kPa}$
$V_t = V_h + V_f - V_p$	$V_f = 1.20 \times 10^{-4} \text{ m}^3$
T is constant	$V_p = ?$

$V_p = \text{volume of powder}$

$P_f = 141 \text{ kPa}$
$V_f = ?$
$V_t = ?$
$V_p = ?$

$$V_t = V_f - V_p$$

$$m_f = 0.038 \text{ kg}$$

T is constant
 $PV = \text{constant}$

$$\begin{aligned}
 P_i V_t &= P_f V_t \\
 110(V_h + V_f - V_p) &= 141(V_f - V_p) \\
 110 V_h + 110 V_f - 110 V_p &= 141 V_f - 141 V_p \\
 31 V_p &= 31 V_f - 110 V_h \\
 31 V_p &= 31(1.2 \times 10^{-4}) - 110(27.27 \times 10^{-6}) \\
 31 V_p &= (3.72 \times 10^{-3}) - (3.00 \times 10^{-3}) \\
 31 V_p &= 0.72 \times 10^{-3} \\
 V_p &= \frac{0.72 \times 10^{-3}}{31} = 2.323 \times 10^{-5} \text{ m}^3 \\
 &= 2.3 \times 10^{-5} \text{ m}^3 \text{ to 2 SF}
 \end{aligned}$$

If $2.7 \times 10^{-5} \text{ m}^3$ is used final answer is 1652 kg m^{-3}

$$\text{Density } \rho = \frac{m_f}{V_p} = \frac{0.038}{2.323 \times 10^{-5}} = 1635 \text{ kg m}^{-3} = 1600 \text{ kg m}^{-3} \text{ to 2 SF}$$

If $V_p = 2.7 \times 10^{-5} \text{ m}^3$ is used final answer is $\rho = 1652 \text{ kg m}^{-3}$

If $V_p = 2.4 \times 10^{-5} \text{ m}^3$ is used final answer is $\rho = 1583 \text{ kg m}^{-3}$

Conceptual Physics Chapter 20 Gases Exercises Answers

D Keegan



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1. Understanding the eBook Conceptual Physics Chapter 20 Gases Exercises Answers
 - The Rise of Digital Reading Conceptual Physics Chapter 20 Gases Exercises Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Conceptual Physics Chapter 20 Gases Exercises Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Conceptual Physics Chapter 20 Gases Exercises Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Conceptual Physics Chapter 20 Gases Exercises Answers
 - Personalized Recommendations
 - Conceptual Physics Chapter 20 Gases Exercises Answers User Reviews and Ratings
 - Conceptual Physics Chapter 20 Gases Exercises Answers and Bestseller Lists
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- 6. Navigating Conceptual Physics Chapter 20 Gases Exercises Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Conceptual Physics Chapter 20 Gases Exercises Answers Compatibility with Devices
 - Conceptual Physics Chapter 20 Gases Exercises Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Conceptual Physics Chapter 20 Gases Exercises Answers
 - Highlighting and Note-Taking Conceptual Physics Chapter 20 Gases Exercises Answers
 - Interactive Elements Conceptual Physics Chapter 20 Gases Exercises Answers
- 8. Staying Engaged with Conceptual Physics Chapter 20 Gases Exercises Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Conceptual Physics Chapter 20 Gases Exercises Answers
- 9. Balancing eBooks and Physical Books Conceptual Physics Chapter 20 Gases Exercises Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Conceptual Physics Chapter 20 Gases Exercises Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Conceptual Physics Chapter 20 Gases Exercises Answers
 - Setting Reading Goals Conceptual Physics Chapter 20 Gases Exercises Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Conceptual Physics Chapter 20 Gases Exercises Answers
 - Fact-Checking eBook Content of Conceptual Physics Chapter 20 Gases Exercises Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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