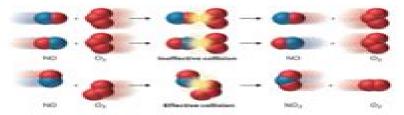
Collision Theory Summary

Main Idea: Particles must collide to react.

However, not all collisions are successful/effective. In other words, not all collisions result in a chemical reaction.

In order for reactants to be converted to products two requirements must be met:

- The collision must have sufficient energy to overcome the required activation energy of the reaction.
- The particles must collide at the correct orientation to allow a reaction to occur.



There are three main ways to increase successful/effective collisions:

- Increase the collision frequency. (Increase how often the particles collide.)
- Increase the collision energy.
- Increase correct collision orientation.

The rate of a chemical reaction is dependent on the rate of successful particle collisions.

Using the above ideas, we can describe five factors (nature of reactants, concentration, addition of a catalyst, exposed surface area, temperature) that affect the rate of a reaction.

Of the five factors which factors will affect each of the following variables:

Collision Frequency	Collision Energy	Collision Orientation
Particle speed affects how often the particles collide Concentration The number of particles in a given space affects how often the particles collide Nature of Reactants The state/phase of the reactants affects their movement and therefore, how often the particles collide Surface Area The amount of exposure the reactant particles have to each other affects how often the particles collide	Temperature Particle speed affects the energy/force the particles collide with	Nature of Reactants Different substances have particles with different sizes, shapes and complexity. Particles that are more complex and non-symmetrical will need to collide in a more specific way (ex. with one end or part of the particle) than simpler particles. Catalysts Catalysts Catalysts provide an alternate way for the reaction to occur, often by orienting the reacting particles in such a way that successful collisions are more likely

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Chris Sunley

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