

EXPERIMENT 2: Analysis of Ions in Commercial Preparations

DATE: September 8, 2023

A. Results

1) Alum

a) Physical Tests:

- i) Color: White
- ii) Physical State: Fine crystalline powder

b) Chemical Tests:

a.) 1. A white precipitate is formed upon adding 1 mL of 1N NaOH.

2. The white precipitate dissolves upon adding excess 1N NaOH but the color of the sample changed to a clear yellowish substance. (Aluminum Ion is present and $\text{Al}(\text{OH})_4^-$ is formed)

b.) 1. In addition of sodium bitartrate to the alum solution there is no precipitate formed.

2. There's still no precipitate formed after adding excess of Sodium Bitartrate but the color of the sample change to a very light greenish substance.

c.) 1. A clear white precipitate is formed upon adding 6N Ammonium Hydroxide.

2. An excessive amount of ammonium hydroxide produced a precipitate that was still clear and white but did not dissolve.

d.) 1. The alum solution formed a milky white cloud precipitate. (Sulfate ion is present)

2.

Ion(s) Present: Aluminum Ion (Al^{+3}), sulfate ion (SO_4^{2-})

2) Baking Soda

a) Physical Tests:

- i) Color: White (powder)
- ii) Physical State: odorless fine powdered solid matter

b) Chemical Tests:

A. N/A

B. Small bubbles gradually rising up with the addition of 1 mL of 1N sulfuric acid; many bubbles quickly rising up along with a white precipitate, in the addition of 1 mL of calcium hydroxide T.S.

C. N/A

c. Ion(s) present: Sodium ion (Na^+) and Bicarbonate ion (HCO_3^-)

3) Laundry Bleaching Agent

a) Physical Tests:

- i) Color: Yellow transparent
- ii) Physical State: Liquid

b) Chemical Test:

- i) Red litmus paper turns to pale blue which signifies the substance as base (alkaline).
- ii) It effervesces, and the color of the combined substance is transparent yellow.

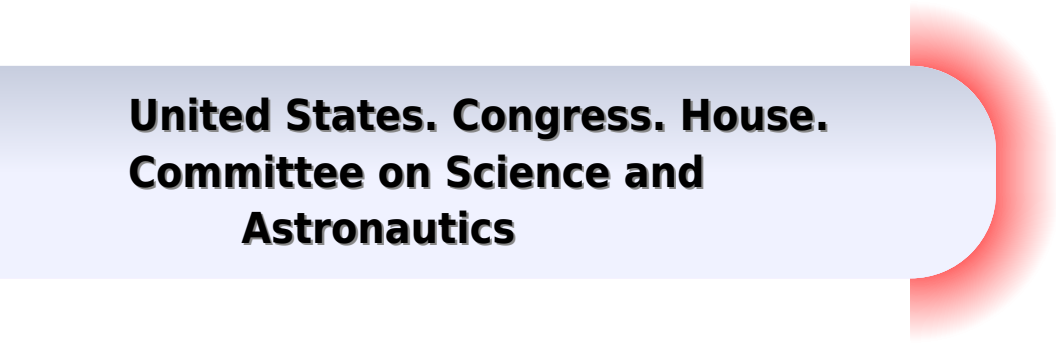
4) Calamine Lotion

a) Physical Tests:

- i) Color: Light cream orange

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