

Worksheet: Neutralization: Chapter 21, (p. 612–630 only)**Page 1**

Complete and *balance* these neutralization reactions. Label the acid (starts with "H"), base (ends in "OH") and *name* the salt produced. Be sure to write the *correct formula* for the salt!

I've named the unfamiliar acids for you. Remember that -ic acids make -ate ions and only the FIRST H's are acidic in organic acids. (p. 613–615)



base

benzoic acid

Salt = magnesium benzoate

**Hydrolysis and Buffers** (p. 626–630)

9. Identify the acid, base, conjugate acid and conjugate base in these hydrolysis reactions.



10. Identify the following as Salts of strong or weak bases and strong or weak acids.
(HCl, HNO₃ and H₂SO₄ are strong acids. NH₃ [NH₄OH] is a weak base.)

Would their solutions in water be acidic, basic or neutral?

weak/strong base

weak/strong acid

acidic, basic or neutral solution?



11. H₂CO₃ is a *much weaker* acid than HC₂H₃O₂. Would solutions of each of these *salts* be:

Neutral, a little Basic or Very Basic? NaCl ____ NaC₂H₃O₂ ____ Na₂CO₃ ____ Explain.

Chemistry Neutralization And Answer Key

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