

COMPOSITION OF HYDRATES

Topic: Hydrates

A hydrate is an ionic compound with water molecules loosely bonded to its crystal structure. The number of water molecules bonded to each unit of the salt. For example, the formula $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ indicates that there are x water molecules for every one formula unit of CuSO_4 . Answer the questions below.

1. What percentage of water is found in $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$?

$$C_{\text{Cu}} = 64$$

$$S_{\text{S}} = 32$$

$$O_{\text{O}} = 16$$

$$H_{\text{H}} = 1 \quad \text{Total} = 159.5$$

$$20 \text{ H}_2\text{O} = \frac{360}{159.5} = 2.26\%$$

$$36.7\%$$

2. What percentage of water is found in $\text{FeSO}_4 \cdot x\text{H}_2\text{O}$?

$$\text{Fe} = 56 \quad S_{\text{S}} = 32 \quad O_{\text{O}} = 16$$

$$H_{\text{H}} = 1 \quad \text{Total} = 159.5$$

$$18 \text{ H}_2\text{O} = \frac{324}{159.5} = 2.03\%$$

$$\frac{18 \times 18}{159.5} \times 100 = 20.3\%$$

$$6.57\%$$

3. A 5.0 gram sample of a hydrate of BaCl_2 was heated, resulting in 3.0 g of the anhydrous salt remaining. What percentage of water was in the hydrate?

$$\text{Ba} = 137 \quad \text{Cl} = 35.5 \quad \text{H}_2\text{O} = 18$$

$$\text{Hydrate} = \text{mass} \times \text{anhydrous}$$

$$\text{Water} = \text{mass} \times \text{anhydrous} \quad \text{Total} = \text{mass} \times \text{anhydrous} \quad \text{Total} = \text{mass} \times \text{anhydrous}$$

$$\frac{5.0 \text{ g} \times 137}{137 + 35.5} = 3.0 \text{ g}$$

$$\frac{2.0 \text{ g}}{5.0 \text{ g}} = 40\%$$

4. A 2.5 g sample of a hydrate of $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ was heated, and only 1.7 g of the anhydrous salt remained. What percentage of water was in the hydrate?

$$\frac{2.5 \text{ g} \times 159.5}{159.5} = 1.7 \text{ g}$$

$$\frac{0.8 \text{ g}}{2.5 \text{ g}} \times 100 = 32\%$$

$$\frac{0.8 \text{ g}}{2.5 \text{ g}} = 32\%$$

5. A 3.0 gram sample of $\text{Na}_2\text{CO}_3 \cdot x\text{H}_2\text{O}$ is heated to constant mass. How much anhydrous salt remains?

$$2 \text{ Na} = 46 \quad \text{C} = 12 \quad \text{O}_3 = 48 \quad \text{H}_2\text{O} = 18$$

$$\frac{18 \text{ g}}{112 \text{ g}} = 16.07\%$$

$$\frac{3.0 \text{ g} \times 112}{112} = 2.67 \text{ g}$$

$$\frac{18 \text{ g}}{112 \text{ g}} = 16.07\%$$

$$2.67 \text{ g}$$

6. A 5.0 gram sample of $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ was heated, and 3.0 g of the anhydrous salt remained. Determine the value of x .

$$\frac{5.0 \text{ g} \times 159.5}{159.5} = 3.0 \text{ g}$$

$$\frac{1.0 \text{ g} \times 159.5}{159.5} = 0.625 \text{ g}$$

$$\frac{1.0 \text{ g} \times 159.5}{159.5} = 0.625 \text{ g}$$

$$\frac{1.0 \text{ g}}{0.625 \text{ g}} = 1.6$$

$$1.6 \times 5 = 8$$

Chemistry Lab 11 Composition Of Hydrates Answers

Karin Nielsen-Saines



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