

Acids and Bases

Overview of Acids and Bases

Indicators: Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to form a bond (LBP)



Remember: (1) only H atoms, (2) only OH is a base

What about organic indicators and how organic indicators?

Acids and Bases

Acids are proton donors. (The tendency to bring in more electrons to bring the lone pair to the base.)



General acid-base reactions of acid-base reactions:

to find the strong base, the base gets protonated, change of the base and usually for weak acids hydrogens of the acid.

to find the strong acid, the acid gets deprotonated, change of the acid and usually for weak bases hydrogens of the acid.

Remember: -OH is a weak base, -OH- is a strong base

1. Acid: Donates its proton to base (proton donor). What if there is no proton transfer?

Acids and Bases

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Acid-base reactions: acids and bases are described in terms of electron.

Acid: Donates its proton to base (proton donor). What if there is no proton transfer?

Acid Strength

What does it mean "strong acid"?

Indicating: the more the acid donates its proton, the stronger it is.



The dissociation of the acid is a reversible reaction that is described by equilibrium constant.



Acid Strength and pKa

The Ka value of an acid (acid) and its conjugate base (B-) are shown below.



The acid is a stronger acid than its conjugate base.

To compare the acid, the pKa is used, which is the pKa of the acid.

$$\text{pKa} = -\log K_a$$

$$\text{pKa}_1 = -\log K_{a1} = 10, \quad \text{pKa}_2 = -\log K_{a2} = 10$$

Acid: the stronger the acid, the lower the pKa value. (It is to separate the Ka)

Acid: the stronger the acid, the lower the pKa value.



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base. The more stable the conjugate base, the stronger the acid.

The stronger the acid, the more stable the conjugate base.



Acid: the stronger the acid, the more stable the conjugate base.

Acid: the stronger the acid, the more stable the conjugate base.

1. The first and last factor is the **charge** to which the conjugate base is converted. Usually, an anion is more stable than a cation.

2. The more stable the conjugate base, the stronger the acid.



1. When comparing different acids, the more stable the conjugate base, the stronger the acid.

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The Position of Equilibrium for an Acid-Base Reaction

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Choosing a Base to Deprotonate a Given Compound

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Choosing an Acid to Protonate a Given Compound

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Hammett-Hansch Equation

Indicators: the more stable the conjugate base, the stronger the acid.

$$\text{pKa} = \text{pKa}_0 + \rho \log \frac{K}{K_0}$$

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Leo J. Malone, Theodore O. Dolter



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 - The Rise of Digital Reading Chemistry Study Guide Acids And Base
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Acids And Base
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Acids And Base
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Acids And Base
 - Personalized Recommendations
 - Chemistry Study Guide Acids And Base User Reviews and Ratings
 - Chemistry Study Guide Acids And Base and Bestseller Lists
5. Accessing Chemistry Study Guide Acids And Base Free and Paid eBooks
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Acids And Base
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Acids And Base
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Chemistry Study Guide Acids And Base
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide Acids And Base
 - Fact-Checking eBook Content of Chemistry Study Guide Acids And Base
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements

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