

CHEMISTRY I HONORS SEMESTER 2 EXAM REVIEW

Unit 1: Stoichiometry

Essential questions:

- How does the amount of each reactant present at the start of a chemical reaction determine how much product is formed?
- How are balanced chemical equations used when solving stoichiometric problems?
- What occurs chemically in a reaction?
- What is a limiting reagent?
- What is the percent yield of a chemical reaction?

Key vocabulary:

- | | | |
|-----------------|---------------------|-----------------|
| • reactant | • chemical reaction | • actual yield |
| • product | • limiting reagent | • percent yield |
| • stoichiometry | • excess reagent | |
| • mole ratio | • theoretical yield | |

Reaction:

1. Differentiate between the significance of the coefficients in a balanced chemical equation and the significance of the subscripts in a chemical formula. _____

2. State the law of Conservation of Mass and explain its relationship to stoichiometry. _____

3. One disadvantage of burning propane (C_3H_8) is that carbon dioxide is one of the products. The released CO_2 increases the growing concentration of CO_2 in the atmosphere. How many moles of carbon dioxide are produced when 30.0 moles of propane are burned in excess oxygen and is gas? _____ $C_3H_8 +$ _____ $O_2 \rightarrow$ _____ $CO_2 +$ _____ H_2O
4. Water decomposes to produce hydrogen gas and oxygen gas. How many grams of water would be required to produce 30.0 moles of hydrogen gas? _____ $H_2O \rightarrow$ _____ $H_2 +$ _____ O_2
5. If 25.0 grams of carbon dioxide are used in photosynthesis how many moles of glucose ($C_6H_{12}O_6$) could be produced according to the following equation: _____ $6CO_2 +$ _____ $6H_2O \rightarrow$ _____ $C_6H_{12}O_6 +$ _____ $6O_2$
6. One series of reactions that inflates air bags in automobiles is the decomposition of sodium azide (NaN_3). Determine the amount of N_2 produced if 100.0 g of NaN_3 is decomposed using the following unbalanced equation: _____ $NaN_3 \rightarrow$ _____ $Na +$ _____ N_2
7. Nitrogen gas reacts with hydrogen gas to produce ammonia: _____ $N_2 +$ _____ $H_2 \rightarrow$ _____ NH_3
 If 100.0 g of nitrogen gas reacts with 30.0 g of hydrogen gas, what mass of ammonia will be produced?

Chemistry Honors Semester 2 Exam Review Answers

CollegeBoard

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