

Assessment #3 Review

Name: _____

You must show all work for credit.

Due Date: _____ (No late packets will be accepted)

I. Balance Equations*Balance the equations*

- $\text{Mg} + \underline{2} \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + \underline{2} \text{NaBr} \rightarrow \text{MgBr}_2 + \underline{2} \text{Na}$
- $\text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \text{Fe} + \underline{3} \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of Fe_2O_3 in reaction #6? 3

Write the balanced equation for problems # 8 and 9:

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on left, 2 on right, so no net created or lost

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

2 N, 6 H, 2 on left, 2 on right

12. Using the balanced equation shown in problem #10, list the products of the reaction.

2NH₃

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

N₂, 3H₂

II. Types of Reactions*Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



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need to have wicked competencies the ability to respond easily and on the fly to complex challenges Unfortunately a traditional education that focuses on content and skills often fails to achieve this sense of wickedness Students memorize for the test prepare for the paper practice the various algorithms over and over again but when the parameters or dynamics of the test or the paper or the equation change students are often at a loss for how to adjust This is a course design book centered on the idea that the goal in the college classroom in all classrooms all the time is to develop students who are not just loaded with content but capable of using that content in thoughtful deliberate ways to make the world a better place Achieving this goal requires a top to bottom reconsideration of courses including student learning goals text selection and course structure day to day pedagogies and assignment and project design Creating Wicked Students takes readers through each step of the process providing multiple examples at each stage while always encouraging instructors to consider concepts and exercises in light of their own courses and students

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