

Corrected
Answered

Dr. David
R. J. Smith

10. (10 points) Given the following reaction, write the balanced equation for the reaction of the reactants with the products (10 points each).

1. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

2. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

3. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

4. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

5. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

6. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

7. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

8. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

9. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

10. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

11. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

12. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

13. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

14. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

15. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

16. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

17. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

18. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

19. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

20. $\text{Pb}^{2+} + \text{Cl}^- \rightarrow \text{PbCl}_2$ (10 points)

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