

Acid/Base Strength of Conjugate Pairs

(How Are K_a and K_b Related?)

WARM-UP

Model 1: The Mathematical Relationship between K_a and K_b of a Conjugate Pair.

Acid	K_a	Conjugate Base	K_b	$K_a \times K_b$
HF hydrofluoric acid	$\frac{[\text{H}_3\text{O}^+][\text{F}^-]}{[\text{HF}]}$	F ⁻ fluoride ion	$\frac{[\text{OH}^-][\text{HF}]}{[\text{F}^-]}$	
HONO nitrous acid				
		NH ₃ ammonia		

Critical Thinking Questions

- Fill in the missing entries in Model 1.
- For each acid and each conjugate base in Model 1, write the balanced chemical reaction that has the K_a or K_b as its equilibrium constant. The conjugate base F⁻ is worked as an example:



END OF WARM-UP

Chemactivity 43 Acid Strength

Mike Jess



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