

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

States of Matter

Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

1. The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
2. The kinetic molecular theory makes the following assumptions:
 - a. In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
 - b. Because gas particles are far apart, there are no significant attractive or repulsive forces between gas particles.
 - c. Gas particles are in constant and random motion.
 - d. The collisions between gas particles are elastic; that is, no kinetic energy is lost.
3. The kinetic energy of a particle is represented by the equation $KE = \frac{1}{2}mv^2$.
4. Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- | | |
|--------------|--|
| <u>True</u> | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas. |
| <u>True</u> | 2. The random motion of gas particles causes a gas to expand until it fills its container. |
| <u>False</u> | 3. The density of a gas decreases as it is compressed. |
| <u>True</u> | 4. A gas can flow into a space occupied by another gas. |
| <u>True</u> | 5. The diffusion of a gas is caused by the random motion of the particles of the gas. |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles. |
| <u>True</u> | 7. During effusion, a gas escapes through a tiny opening into a vacuum. |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

Section 13.2 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

13. Pressure is defined as force per unit c. area.
14. What is an instrument designed to measure atmospheric pressure? c. sphygmomanometer
15. The height of the liquid in a barometer is affected by all of the following EXCEPT the a. altitude.
16. The pressure of the gas in a manometer is directly related to which of the following quantities? b. height of the mercury column in the closed end arm
17. One atmosphere is equal to a pressure of b. 101.3 kPa.
18. The partial pressure of a gas depends on all of the following EXCEPT the b. size of the container.
19. The pressure of a sample of air in a manometer is 102.1 kPa. What is the partial pressure of nitrogen (N_2) in the sample if the combined partial pressures of the other gases is 22.4 kPa? b. 79.7 kPa

Use the figure to answer the following questions.

20. What instrument is illustrated in the figure? barometer
21. Who invented this instrument? Evangelista Torricelli
22. What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
23. What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



Chemistry Study Guide Answers States Of Matter

Arshad Iqbal



Chemistry Study Guide Answers States Of Matter:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Table of Contents Chemistry Study Guide Answers States Of Matter

1. Understanding the eBook Chemistry Study Guide Answers States Of Matter
 - The Rise of Digital Reading Chemistry Study Guide Answers States Of Matter
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Answers States Of Matter
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Answers States Of Matter
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Answers States Of Matter
 - Personalized Recommendations
 - Chemistry Study Guide Answers States Of Matter User Reviews and Ratings

- Chemistry Study Guide Answers States Of Matter and Bestseller Lists
- 5. Accessing Chemistry Study Guide Answers States Of Matter Free and Paid eBooks
 - Chemistry Study Guide Answers States Of Matter Public Domain eBooks
 - Chemistry Study Guide Answers States Of Matter eBook Subscription Services
 - Chemistry Study Guide Answers States Of Matter Budget-Friendly Options
- 6. Navigating Chemistry Study Guide Answers States Of Matter eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Study Guide Answers States Of Matter Compatibility with Devices
 - Chemistry Study Guide Answers States Of Matter Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide Answers States Of Matter
 - Highlighting and Note-Taking Chemistry Study Guide Answers States Of Matter
 - Interactive Elements Chemistry Study Guide Answers States Of Matter
- 8. Staying Engaged with Chemistry Study Guide Answers States Of Matter
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Answers States Of Matter
- 9. Balancing eBooks and Physical Books Chemistry Study Guide Answers States Of Matter
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Answers States Of Matter
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Study Guide Answers States Of Matter
 - Setting Reading Goals Chemistry Study Guide Answers States Of Matter
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Study Guide Answers States Of Matter
 - Fact-Checking eBook Content of Chemistry Study Guide Answers States Of Matter
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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