

pH AND pOH

Chapter 14 Section 14.2

The pH of a solution indicates how acidic or basic that solution is.

pH ranges of 0-7 = acidic

7 = neutral

7-14 = basic

Since $\text{pH} + \text{pOH} = 14$ at 25°C, if you know the pH of a solution, then you can calculate the pOH of that solution.

$\text{pH} = 4 \Rightarrow \text{pOH} = 10$

So if $\text{pH} = 10$ then $\text{pOH} = 4$.

$\text{pOH} = 4 \Rightarrow \text{pH} = 10$

So if $\text{pOH} = 10$ then $\text{pH} = 4$.

Together $\text{pH} + \text{pOH} = 14$.

Complete the following chart.

	pH	pH	pOH	pOH	Acids or Bases
1.	normal	8	normal	6	Acidic
2.	10^{-10} M	10	10^{-4}	4	acidic
3.	10^{-10} M	10	normal	4	acidic
4.	normal	12	10^{-12} M	12	acidic
5.	10^{-12} M	12	10^{-2} M	2	acidic
6.	10^{-12} M	12	10^{-2} M	2	acidic
7.	10^{-7} M	7	normal	7	neutral
8.	normal	14	10^{-14} M	14	acidic
9.	10^{-8} M	8	10^{-6} M	6	acidic
10.	10^{-8} M	6	10^{-6} M	6	acidic

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Daniel F McAuley



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