

**CHEMFILE MINI-GUIDE TO PROBLEM SOLVING****General Plan for Solving Dilution Problems**

**1**  $M_1V_1 = M_2V_2$

Rearrange the equation  $M_1V_1 = M_2V_2$  algebraically to solve for the unknown quantity.

**2**

The equation used to calculate the unknown quantity will be one of the following four:

$$V_2 = \frac{M_1V_1}{M_2}, M_2 = \frac{M_1V_1}{V_2}, V_1 = \frac{M_2V_2}{M_1}, M_1 = \frac{M_2V_2}{V_1}$$

Substitute each of the known quantities for its symbol, and calculate.

**3**

Unknown molarity or volume

**SAMPLE PROBLEM 1**

What is the molarity of a solution that is made by diluting 50.00 mL of a 4.74 M solution of HCl to 250.00 mL?

**SOLUTION****1. ANALYZE**

- *What is given in the problem?* the molarity of the stock solution, the volume used to dilute, and the volume of the diluted solution
- *What are you asked to find?* the molarity of the diluted solution

| Items                                           | Data       |
|-------------------------------------------------|------------|
| Concentration of the stock solution ( $M_1$ )   | 4.74 M HCl |
| Volume of stock solution used ( $V_1$ )         | 50.00 mL   |
| Volume of diluted solution ( $V_2$ )            | 250.00 mL  |
| Concentration of the diluted solution ( $M_2$ ) | ? M        |

# Chemfile Mini Guide To Problem Solving Answer Key

**Rachel Turoscy**



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