

# CHAPTER 14 STUDY GUIDE

## Mixtures and Solutions

### Section 14.1 Heterogeneous and Homogeneous Mixtures

*In your textbook, read about suspensions and colloids.*

For each statement below, write *true* or *false*.

- |              |                                                                                                                                   |
|--------------|-----------------------------------------------------------------------------------------------------------------------------------|
| <u>false</u> | 1. A solution is a mixture containing particles that settle out of the mixture if left undisturbed.                               |
| <u>true</u>  | 2. The most abundant substance in a colloid is the dispersion medium.                                                             |
| <u>false</u> | 3. A colloid can be separated by filtration.                                                                                      |
| <u>true</u>  | 4. A solid emulsion consists of a liquid dispersed in a solid.                                                                    |
| <u>true</u>  | 5. Whipped cream is an example of a foam.                                                                                         |
| <u>false</u> | 6. In an aerosol, the dispersing medium is a liquid.                                                                              |
| <u>true</u>  | 7. Brownian motion results from the collisions of particles of the dispersion medium with the dispersed particles.                |
| <u>true</u>  | 8. Dispersed particles in a colloid do not tend to settle out because they have polar or charged atomic groups on their surfaces. |
| <u>false</u> | 9. Stirring an electrolyte into a colloid stabilizes the colloid.                                                                 |
| <u>true</u>  | 10. Colloids demonstrate the Tyndall effect.                                                                                      |

The table below lists the characteristics of particles in colloids, solutions, and suspensions. Place a check in the column of each mixture whose particles have a particular characteristic.

| Characteristics of Particles             | Colloid                             | Solution                            | Suspension                          |
|------------------------------------------|-------------------------------------|-------------------------------------|-------------------------------------|
| 11. Less than 1 nm in diameter           |                                     | <input checked="" type="checkbox"/> | <input checked="" type="checkbox"/> |
| 12. Between 1 nm and 1000 nm in diameter | <input checked="" type="checkbox"/> | <input checked="" type="checkbox"/> | <input checked="" type="checkbox"/> |
| 13. More than 1000 nm in diameter        | <input checked="" type="checkbox"/> | <input checked="" type="checkbox"/> |                                     |
| 14. Settle out if undisturbed            |                                     |                                     |                                     |
| 15. Pass through standard filter paper   | <input checked="" type="checkbox"/> |                                     |                                     |
| 16. Lower vapor pressure                 |                                     |                                     | <input checked="" type="checkbox"/> |
| 17. Scatter light                        | <input checked="" type="checkbox"/> |                                     |                                     |

# Chem Mixtures And Solutions Study Guide Answers

**James E. Brady, Fred Senese**



## **Chem Mixtures And Solutions Study Guide Answers:**

*Organic Chemistry Study Guide with Solutions Manual* Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer *Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement [Study Guide to Accompany Basics for Chemistry](#) Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to

learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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