

Collision Theory Summary

Main Idea: Particles must collide to react.

However, not all collisions are successful/effective. In other words, not all collisions result in a chemical reaction.

In order for reactants to be converted to products two requirements must be met:

1. The collision must have sufficient energy to overcome the required activation energy of the reaction.
2. The particles must collide at the correct orientation to allow a reaction to occur.



There are three main ways to increase successful/effective collisions:

1. Increase the collision frequency. (Increase how often the particles collide.)
2. Increase the collision energy.
3. Increase correct collision orientation.

The rate of a chemical reaction is dependent on the rate of successful particle collisions.

Using the above ideas, we can describe five factors (nature of reactants, concentration, addition of a catalyst, exposed surface area, temperature) that affect the rate of a reaction.

Of the five factors which factors will affect each of the following variables:

Which factor(s) affects the ...		
Collision Frequency	Collision Energy	Collision Orientation
Temperature Particle speed affects how often the particles collide Concentration The number of particles in a given space affects how often the particles collide Nature of Reactants The state/phase of the reactants affects their movement and therefore, how often the particles collide Surface Area The amount of exposure the reactant particles have to each other affects how often the particles collide	Temperature Particle speed affects the energy/force the particles collide with	Nature of Reactants Different substances have particles with different sizes, shapes and complexity. Particles that are more complex and non-symmetrical will need to collide in a more specific way (ex. with one end or part of the particle) than simpler particles. Catalysts Catalysts provide an alternate way for the reaction to occur, often by orienting the reacting particles in such a way that successful collisions are more likely

Collision Theory Impact For A Chemical Reaction Answers

Rosina Ehmann



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Collision Theory and Statistical Theory of Chemical Reactions S. G. Christov, 2012-12-06 Since the discovery of quantum mechanics more than fifty years ago the theory of chemical reactivity has taken the first steps of its development The knowledge of the electronic structure and the properties of atoms and molecules is the basis for an understanding of their interactions in the elementary act of any chemical process The increasing information in this field during the last decades has stimulated the elaboration of the methods for evaluating the potential energy of the reacting systems as well as the creation of new methods for calculation of reaction probabilities or cross sections and rate constants An exact solution to these fundamental problems of theoretical chemistry based on quantum mechanics and statistical physics however is still impossible even for the simplest chemical reactions Therefore different approximations have to be used in order to simplify one or the other side of the problem At present the basic approach in the theory of chemical reactivity consists in separating the motions of electrons and nuclei by making use of the Born Oppenheimer adiabatic approximation to obtain electronic energy as an effective potential for nuclear motion If the potential energy surface is known one can calculate in principle the reaction probability for any given initial state of the system The reaction rate is then obtained as an average of the reaction probabilities over all possible initial states of the reacting particles In the different stages of this calculational scheme additional approximations are usually introduced

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estimates of kinetic parameters have been included in this version of the book Finally kinetics requires knowledge that must be complemented and tested in the laboratory Therefore practical examples of reactions performed in bench and semi pilot scales are discussed in the final chapter This edition of the book has been organized in two parts In the first part a thorough discussion regarding reaction kinetics is presented In the second part basic equations are derived and used to represent the performances of batch and continuous ideal reactors isothermal and non isothermal reaction systems and homogeneous and heterogeneous reactor vessels as illustrated with several examples and exercises This textbook will be of great value to undergraduate and graduate students in chemical engineering as well as to graduate students in and researchers of kinetics and catalysis

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