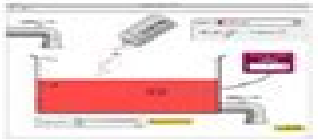
Concentration and Molarity PhET Labs

Name: _____

Procedure: https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html



Concentration

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all the sliders. Notice what happens to the concentration as solid solute is added and when evaporation occurs.

1. How does the concentration change as solid solute is added? concentration increases
2. How does the concentration change as additional water is added? concentration decreases
3. How does the concentration change as evaporation occurs? concentration increased
4. How do you know when a solution is saturated? concentration does not change (it also shows on screen)
6. When a solution is saturated, and additional solid solute is added, what happens? solid does not dissolve and settles at bottom
7. Why do you think this is? solvent cannot hold any more solute
8. How does adding this additional solute change the concentration of this saturated solution? decreases concentration
9. How does evaporation change the concentration of a saturated solution? no change- solution is already saturated

Part 2: Concentrated Solutions (some website as part 1 just click on the Solution option)

10. Adding a concentrated solution … describe a way to determine the concentration of the solution in the spigot. Write your plan here: empty the container and add only one solution from the dropper and then use the sensor to measure concentration

11. Using your plan…how might you get that concentrated solution to become saturated? you can evaporate

12. Does your plan work for all the other solutions too? yes Why? / Why Not? evaporation reduces the amount of water (or solute) so concentration increases

Part 3: Molarity https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html

$$\text{Molarity (M)} = \frac{\text{amount of solute (mol)}}{\text{volume of solution (L)}}$$

Molarity is moles per Liter, that is, how many moles of solute (entire salt) is dissolved per Liter of solution.

First, determine the saturation concentration of each of the solutions, that is, how concentrated can you get each solution before the solution is saturated. If you can't determine the concentration using the simulation "Molarity", try using the simulation "Concentration" (You will use this information again in *Part 5*, if your instructor requires it)

Molarity

Cobalt (II) nitrate	5.640 M	Potassium chromate	3.350 M
Cobalt chloride	4.330 M	Nickel (II) chloride	5.210 M
Potassium dichromate	0.510 M	Copper sulfate	1.380 M
Gold (III) chloride	2.30 M	Potassium permanganate	0.480 M

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