

Concentration and Molarity PhET-Chemistry Labs



Introduction: Everyone likes candy. Have you ever wondered how that candy is produced? How do they get all that delicious sugar into those tiny packages? Could you make hard candy like those you can buy? It's easier than you think. Web searching for "rock candy" will yield a number of delicious recipes you can try at home.



Concentration

Some handy vocabulary for you to define:

Solute: a substance that is dispersed in a solution

Solvent: a substance doing the dissolving, usually present in greatest quantity

Moles: the amount of a chemical substance that contains 6.02×10^{23} formula units of the substance.

Molarity: The concentration of a solution in moles of solute per liter of solution

Saturated (not fats): the point at which a solution of a substance can dissolve no more of that substance

Unsaturated (not fats): the point at which a solution of a substance can still dissolve more of that substance

Supersaturated: containing more solute than a saturated solution and therefore not in equilibrium

Procedure: PhET → Play with the Sims → Chemistry → Concentration

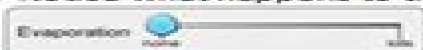
Run Now!

Concentration
(mol/L)

0.179

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all the sliders. Notice what happens to the concentration as solid solute is added and when evaporation occurs.



How does the concentration change as solid solute is added? As solid solute is added, the concentration goes up.

How does the concentration change as additional water is added? As additional water is added, the concentration decreases.

How does the concentration change as evaporation occurs? As evaporation occurs, the concentration goes up.



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Answers Key

Hussin A.Rothana



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