

Chemistry Study Guide

Periodic Table of the Elements

H 1 1.008																	He 2 4.003	
Li 3 6.941	Be 4 9.012											B 5 10.81	C 6 12.01	N 7 14.01	O 8 16.00	F 9 18.99	Ne 10 20.18	
Na 11 22.99	Mg 12 24.31									Al 13 26.98	Si 14 28.09	P 15 30.97	S 16 32.07	Cl 17 35.45	Ar 18 39.95			
K 19 39.10	Ca 20 40.08	Sc 21 44.96	Ti 22 47.88	V 23 50.94	Cr 24 52.00	Mn 25 54.94	Fe 26 55.85	Co 27 58.93	Ni 28 58.71	Cu 29 63.55	Zn 30 65.38	Ga 31 69.72	Ge 32 72.64	As 33 74.92	Se 34 78.96	Br 35 79.90	Kr 36 83.80	
Rb 37 85.47	Sr 38 87.62	Y 39 88.91	Zr 40 91.22	Nb 41 92.91	Mo 42 95.94	Tc 43 98.00	Ru 44 101.07	Rh 45 102.91	Pd 46 106.42	Ag 47 107.87	Cd 48 112.41	In 49 114.82	Sn 50 118.71	Sb 51 121.76	Te 52 127.60	I 53 126.91	Xe 54 131.29	
Cs 55 132.91	Ba 56 137.33			Hf 72 178.49	Ta 73 180.95	W 74 183.84	Re 75 186.21	Os 76 190.23	Ir 77 192.22	Pt 78 195.08	Au 79 196.97	Hg 80 200.59	Tl 81 204.38	Pb 82 207.2	Bi 83 208.98	Po 84 209	At 85 210	Rn 86 222
Fr 87 223	Ra 88 226			Rf 104 261	Db 105 262	Sg 106 266	Bh 107 264	Hs 108 277	Mt 109 268	Ds 110 271	Rg 111 272	Cn 112 285	Uut 113 284	Fl 114 289	Uup 115 288	Lv 116 293	Uus 117 286	Uuo 118 294
		Lanthanide series LaCePrNdPmSmEuGdTbDyHoErTmYbLu																
		Actinide series AcThPaUaNpPuAmCmBkCfEsFmMdNoLr																

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

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