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# C is for Clown



A CIRCUS OF "C" WORDS  
by Stan and Jan Berenstain

# C Is For Clown Bright And Early S For Beginning Beginners

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Bibliography**



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one cubic decimetre molality molality  $m$  is defined as the number of moles of the solute per kilogram  $kg$  of the solvent  
formula for molality  $\text{molality} = \frac{\text{moles of solute}}{\text{mass of solvent in kg}}$  example of molality

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 $m = \frac{\text{Ca}^{2+}}{\text{OH}^-}$

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$$m = \frac{n}{V} \times \frac{1000 \text{ mL}}{1000 \text{ dL}}$$

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web 1 molarity  $m = \frac{m}{M}$  mass molar mass  $\times 0.4500 \text{ L} = 0.825 \text{ g}$  141.9579 g/mol  $\times 0.0129 \text{ M} = 2$  molality  $0.825 \text{ g} / 141.9579 \text{ g/mol} = 0.00581158 \text{ mol}$  0.00581158 mol / 0.4500 kg = 0.0129 m 3 mole fraction  $n_2 = \frac{0.825 \text{ g}}{141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$   $n_1 = \frac{450.0 \text{ g}}{18.015 \text{ g/mol}} = 24.97918401 \text{ mol}$

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