

INTEGRATED Math 1

Common Core
State Standards

High
School

Algebraic Expressions, Writing Equations, Solving Inequalities, Functions, Arithmetic Sequences, Systems of Equations, and more!

ALGEBRAIC EXPRESSIONS & PROPERTIES ASSIGNMENT

A **variable** is a symbol that is used to represent an unknown quantity. An **algebraic expression** is a mathematical expression that contains one or more variables, one or more numbers, and one or more arithmetic operations.

EX: x
 $5 - 8$
 $4x + 2y + 7$
 $3x^2 + 7x - 1$

Order of Operations

The **order of operations** is a set of rules that determine which operations to perform first when evaluating a numerical expression.

1. Evaluate within grouping symbols.
2. Evaluate all powers.
3. Multiply and divide from left to right.
4. Add and subtract from left to right.

Properties of Real Numbers

For real numbers a , b , and c , the following properties are true.

Property	Expression
Commutative Property of Addition	$a + b = b + a$
Commutative Property of Multiplication	$a \cdot b = b \cdot a$
Associative Property of Addition	$a + (b + c) = (a + b) + c$
Associative Property of Multiplication	$a \cdot (b \cdot c) = (a \cdot b) \cdot c$

Distributive Property	$a \cdot (b + c) = a \cdot b + a \cdot c$
Additive Identity Property	$a + 0 = a$
Additive Inverse Property	$a + (-a) = 0$
Multiplicative Inverse Property	$a \cdot \left(\frac{1}{a}\right) = 1, a \neq 0$
Multiplicative Property of Zero	$a \cdot 0 = 0$

Parts of an Expression

A **term** is a number, variable, or the product of numbers and variables in an expression. **Like terms** are terms that have the same variables raised to the same powers. Two quantities that are being multiplied in an expression are called **factors**. The **coefficient** of a term is the numerical factor of the term. A term that consists of a number and no variables is called a **constant**.

REASONING WITH QUANTITIES

The units of a quantity can help you understand and solve multi-step problems.

Dimensional Analysis

Dimensional analysis is the process of using conversion factors to change from one set of units to another set of units.

EX: Jamie jogged 495 feet in 45 seconds. What was her speed in miles per hour?

Write the given information as a ratio. Then use appropriate conversion factors to change units.

$$\frac{495 \cancel{ft}}{45 \cancel{s}} \cdot \frac{1 \text{ mi}}{5280 \cancel{ft}} = \frac{1 \text{ mi}}{480 \cancel{s}}$$

Convert from feet to miles.

$$\frac{1 \text{ mi}}{480 \cancel{s}} \cdot \frac{60 \cancel{s}}{1 \text{ min}} = \frac{1 \text{ mi}}{8 \text{ min}}$$

Convert from seconds to minutes.

$$\frac{1 \text{ mi}}{8 \cancel{\text{ min}}} \cdot \frac{60 \cancel{\text{ min}}}{1 \text{ hr}} = \frac{7.5 \text{ mi}}{1 \text{ hr}}$$

Convert from minutes to hours.

Jamie's speed was 7.5 miles per hour.

Modeling with Appropriate Units

When modeling with quantities, it is important to choose appropriate units.

EX: Suppose a tourist buys a map of Florida while on vacation. What might the scale of the map be? What units would represent the actual distance between cities?

The map is a small-scale drawing of the state. Distances on the map will be very small compared to the actual distances between cities. A scale for the map might be 1 cm = 250 km or 2 in. = 175 mi. The distances between the actual cities would be measured in miles or kilometers.

Accuracy & Precision

Accuracy describes how close a measured value is to the actual value.

EX: Thomas is 5 feet 10.25 inches tall. A measurement of 5 feet 10 inches is more accurate than a measurement of 5 feet when describing his height.

Precision describes the exactness of a measurement and is dependent on the unit of measure. The **precision unit** is the smallest unit on the tool being used to measure. As the units of measurement get smaller, the measurement gets more precise.

EX: A ruler that measures in millimeters is more precise than a ruler that measures in centimeters.

EX: For a scale that gives weight in pounds and ounces, the precision unit is ounces.

WRITING EQUATIONS

An **equation** is a number sentence that states that two quantities are equal to each other. The verbal expressions **equals**, **is**, **equal to**, and **is the same as** indicate an equation.

Modeling with Equations

Equations can be used to model real-world situations and solve problems.

EX: Reggie's little sister is 3 years younger than him. The sum of their ages is 27 years. How old is Reggie?

- Define a variable.
- Let r represent Reggie's age.
- Translate the verbal description into an equation.

Reggie's age: r
Reggie's sister's age: $r - 3$
Sum of their ages: $r + (r - 3)$
Because the sum of their ages is 27 years, the equation $r + (r - 3) = 27$ models the situation.

You can verify that the solution to the equation is $r = 15$. So, Reggie is 15 years old. In the next section, you will learn several methods for solving equations.

SOLVING EQUATIONS

You can use the following properties of equality to solve equations.

- **Addition Property of Equality:** If you add the same number to both sides of an equation, the two sides remain equal.
- **Subtraction Property of Equality:** If you subtract the same number from both sides of an equation, the two sides remain equal.
- **Multiplication Property of Equality:** If you multiply both sides of an equation by the same number, the two sides remain equal.
- **Division Property of Equality:** If you divide both sides of an equation by the same nonzero number, the two sides remain equal.

Solving Multi-step Equations

Equations that have multiple operations and more complicated expressions are called **multi-step equations**. To solve these types of equations, simplify both sides as much as you can and use the properties you have learned to undo the operations.

EX: A coin-operated car wash costs \$1.75 plus \$0.50 per minute of use. Shavona spent \$7.75 using the car wash. Solve the equation $1.75 + 0.5m = 7.75$ to find the number of minutes, m , that she used the car wash.

$$1.75 + 0.5m = 7.75$$

$$1.75 - 1.75 + 0.5m = 7.75 - 1.75$$

$$0.5m = 6$$

$$2(0.5m) = 2(6)$$

$$m = 12$$

Shavona used the car wash for 12 minutes.

Original equation
Subtract 1.75 from each side.
Simplify.
Multiply each side by 2.
Simplify.

Math Practice 3: Construct viable arguments and critique the reasoning of others.

Each time you apply a property of real numbers or a property of equality to transform an equation, you can justify that both sides of the equation are still equal. As you solve equations, think about which property allows you to transform it into an equivalent form.

Common Core Math 9th Grade

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Common Core Math 9th Grade:

Decoding **Common Core Math 9th Grade**: Revealing the Captivating Potential of Verbal Expression

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$\text{HNO}_3 + \text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{HC}_2\text{H}_3\text{O}_2$ chemical equation - Aug 11 2023

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web when balancing chemical equations our goal is to have the same number of each type of atom on both sides of the equation only change the coefficients these are the numbers in front substances never change the subscripts the small numbers after elements

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web mar 4 2018 $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$ on the left side there are Mg 1 magnesium in elemental form H_2 there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens N_2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens O_6

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web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate MgCO_3 and two moles of nitric acid HNO_3 react to form one mole of magnesium nitrate $\text{Mg}(\text{NO}_3)_2$ one mole of water H_2O and one mole of carbon dioxide CO_2 show chemical structure image reactants

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web magnesium is already balanced on the reactant and product side balancing nitrogen by multiplying HNO_3 by 2 we get $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$ magnesium 2 nitric acid 2 magnesium nitrate 2 hydrogen gas oxygen and hydrogen get automatically balanced by balancing nitrogen

$\text{Mg}(\text{NO}_3)_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + \text{HNO}_3$ chemical equation - Apr 07 2023

web magnesium nitrate sulfuric acid magnesium sulfate nitric acid

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web the balanced net ionic equation for the reaction of magnesium carbonate with nitric acid is a $\text{MgCO}_3 \text{ s} + 2\text{H}^+ \text{ aq} \rightarrow \text{Mg}^{2+} \text{ aq} + \text{CO}_2 \text{ g} + \text{H}_2\text{O} \text{ l}$ b $\text{Mg}^{2+} \text{ aq} + \text{CO}_3^{2-} \text{ aq} + 2\text{H}^+ \text{ aq} \rightarrow 2\text{NO}_3^- \text{ aq} + \text{Mg}$ write the balanced equation when nitric acid reacts with calcium hydroxide to produce water and calcium nitrate

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web magnesium oxide nitric acid gives rise to a magnesium phosphate hydrogen balanced equation on may 27 2019 how do you balance magnesium and nitric acid you ll need to count all of atoms on both directions of the chemical equation to balance $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg(NO}_3)_2 + \text{H}_2\text{O}$

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web how to balance $\text{HNO}_3 + \text{Mg} \rightarrow \text{Mg(NO}_3)_2 + \text{H}_2$ word equation nitric acid magnesium magnesium nitrate hydrogen gas type of chemical reaction for this reaction we have a single displacement reaction balancing strategies in this single displacement reaction the mg replaces the h in the HNO_3

how to balance $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg(NO}_3)_2 + \text{NO}_2 + \text{H}_2\text{O}$ magnesium youtube - May 08 2023

web nov 20 2020 in this video we ll balance the equation $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg(NO}_3)_2 + \text{NO}_2 + \text{H}_2\text{O}$ and provide the correct coefficients for each compound to balance $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg(NO}_3)_2 + \text{NO}_2 + \text{H}_2\text{O}$ you ll need to be

lecture 13 acid base reactions widener university - Dec 03 2022

web magnesium hydroxide nitric acid balanced total equation $\text{Mg(OH)}_2 \text{ aq} + 2\text{HNO}_3 \text{ aq} \rightarrow \text{Mg(NO}_3)_2 \text{ aq} + 2\text{H}_2\text{O}$ total ionic equation $\text{Mg}^{2+} \text{ aq} + 2\text{OH}^- \text{ aq} + 2\text{H}^+ \text{ aq} + 2\text{NO}_3^- \text{ aq} \rightarrow \text{Mg}^{2+} \text{ aq} + 2\text{NO}_3^- \text{ aq} + 2\text{H}_2\text{O}$ net ionic equation $2\text{OH}^- \text{ aq} + 2\text{H}^+ \text{ aq} \rightarrow 2\text{H}_2\text{O}$ sulfuric acid and lithium hydroxide balanced total equation $\text{H}_2\text{SO}_4 \text{ aq} + 2\text{LiOH} \text{ aq} \rightarrow \text{H}_2\text{O}$