

Experiment 5

Colligative Properties of Solutions

Objectives

Colligative properties are those properties of a solution that depend on the number of solute particles in the solution. This experiment will demonstrate how colligative properties can be used to determine the molar mass of a substance.

Background

When a solute is dissolved in a solvent, the properties of the solution are changed by the presence of the solute. The changes in the properties of the solution are called colligative properties. These properties include boiling point, freezing point, vapor pressure, and osmotic pressure. The changes in these properties are called colligative properties because they depend on the number of solute particles in the solution, not on the identity of the solute.

For example, the boiling point of a solution is higher than the boiling point of the pure solvent. This is because the solute particles in the solution interfere with the solvent molecules' ability to escape the liquid phase. The same effect is observed for the freezing point of a solution. The freezing point of a solution is lower than the freezing point of the pure solvent. This is because the solute particles in the solution interfere with the solvent molecules' ability to form a solid phase. The same effect is observed for the vapor pressure of a solution. The vapor pressure of a solution is lower than the vapor pressure of the pure solvent. This is because the solute particles in the solution interfere with the solvent molecules' ability to escape the liquid phase.

In this experiment, you will determine the boiling point of a pure solvent (distilled water) and a solution of a known solute (sucrose). You will also determine the boiling point of a solution of an unknown solute (a mixture of two different substances). The boiling point of a solution is a colligative property because it depends on the number of solute particles in the solution, not on the identity of the solute.

Colligative Properties Catalyst Lab Manual Answers

Michael Seilmaier



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Publisher's Monthly, 1990 *High School Chemistry Teachers Magazine*, 1973 *Selected Water Resources Abstracts*, 1987-12 Handling and Preservation of Fruits and Vegetables by Combined Methods for Rural Areas Gustavo V. Barbosa-Cánovas, Food and Agriculture Organization of the United Nations, 2003-01-01 This manual contains basic information on post harvest handling and marketing operations and storage of fresh and processed fruit and vegetables It includes practical examples of preservation techniques and highlights technological aspects which can prevent biochemical and physicochemical reactions and microbial growth the main causes of quality losses in fruits and vegetables The suggested methodologies combine technologies such as mild heat treatment water activity reduction lowering of the pH and use of anti microbial substances These relatively new technologies have been successfully applied to various tropical and non tropical fruits in different countries of Latin America and are recommended for use in other fruit producing countries around the world **Uniform Trade List Annual**, 1977 The Success Manual for General Chemistry Elizabeth Kean, Catherine Middlecamp, 1986 **The Software Encyclopedia 2000** Bowker Editorial Staff, 2000-05 **Chemical Products and Aerosol News**, 1955 Vol 25 no 3 v 26 Mar 1962 1963 includes the section Aerosol news v 1 2 no 10 *Catalyst: Lab Manual for Chemistry 30a*, 2010

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