

Chemistry Final Exam Study Guide

STRATEGY: Start by reading through your notes to refresh your memory on these topics. Then, use this review sheet as a starting point to identify the areas on which you need to spend more study time. For those areas, go back to homework assignments, quizzes, and reviews to practice more problems. Don't just answer the questions you know! If you don't know an answer, look it up in your notes, book or online. Also, these are not the questions that will be on the test, you need to be able to understand WHY an answer is correct, and not just WHAT the answer is...copying won't help you on the Final Exam!

Introduction Unit (Ch. 1-4)

1. In a lab, the average measured density for Pre-1982 pennies was 7.98 g/cm³. Given that the literature value for the density is 8.92 g/cm³, calculate the percent error.
2. Convert the following numbers into or out of scientific notation.
a. 548,000 b. 0.0000770 c. 1.200×10^3 d. 9.25×10^6
3. Osmium is the densest element with a density of 22.57 g/cm³. Find the mass of a 56.2 cm³ sample of osmium.
4. Perform the following SI prefix conversions.
a. 65.2 mm = ? cm b. 2.3 kg = ? g c. 65,000 ml = ? L d. 0.502 km = ? cm
5. How many milliliters are in a 2.0 quart jug of milk? (1L = 1.06qts)
6. Ms. Townsend spent last weekend grading tests. If she spent 4 min on each test, how many hours did it take her to grade all 74 tests?
7. Calculate the density from the slope of a "Mass vs. Volume" graph.
8. Write the isotope symbol, including atomic number & mass number, for the following isotopes.
a. carbon-14 b. chromium-53 c. nickel-63 d. zirconium-92
9. Complete the table for the following isotopes.

Symbol	Atomic #	Mass#	Protons	Neutrons	electrons
Zn		65			
	20			21	
		74	34		
		40			18

10. Calculate the average atomic mass of copper if 69.17% of the copper atoms occurring in nature are ⁶³Cu and 30.83% are ⁶⁵Cu.

Atoms/Electrons (Ch. 5)

11. Draw an orbital diagram, complete electron configuration and noble gas notation for
a. Na b. C c. Mo d. Se
12. How does an electron become excited? What does it do when it returns to the ground state?
13. What is a photon? Quantum?
14. Describe the relationship between wavelength and frequency.
15. What type of electromagnetic radiation is emitted when an electron moves from n=4 to n=1?
16. Where are the s,p,d,f blocks located on the periodic table? How many orbitals are associated with each? How many electrons can be in each orbital?

Periodic Table (Ch. 6)

17. DEFINE each of the following trends, describe WHAT the trend is (increases up, etc), explain WHY each trend occurs
a. Atomic Radius b. Ionic Radius
c. Ionization Energy d. Electronegativity
18. Circle the atom with the LARGER radius.
a. Be N b. Ne Xe
19. Circle the particle with the LARGER radius.
a. Cl Cl⁻ b. Mg Mg²⁺
20. Circle the atom with the HIGHER first ionization energy.
a. Li Cs b. Ca As
21. Circle the atom with the HIGHER electronegativity.
a. Cl Si b. O Po

College Chemistry Final Exam Study Guide

Charles Dickens



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