

Concentration and Molarity PhET Labs

Name: _____

Procedure: https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all the sliders. Notice what happens to the concentration as solid solute is added and when evaporation occurs.



Concentration

- How does the concentration change as solid solute is added? concentration increases
 - How does the concentration change as additional water is added? concentration decreases
 - How does the concentration change as evaporation occurs? concentration increases
 - How do you know when a solution is saturated? concentration does not change (it also shows on screen)
 - When a solution is saturated, and additional solid solute is added, what happens? solid does not dissolve and settles at bottom
 - Why do you think this is? solvent can't hold any more solute
 - How does adding this additional solute change the concentration of this saturated solution? decreases concentration
 - How does evaporation change the concentration of a saturated solution? no change- solution is already saturated
- Part 2: Concentrated Solutions (same website as part 1 just click on the Solution option)**
- Adding a concentrated solution ... describe a way to determine the concentration of the solution in the spigot. Write your plan here: empty the container and add only one solution from the dropper and then use the sensor to measure concentration



- Using your plan ... how might you get that concentrated solution to become saturated? you can evaporate
- Does your plan work for all the other solutions too? yes Why? / Why Not? evaporation reduces the amount of water (or solute) so concentration increases



Part 3: Molarity https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html

Molarity (M) = $\frac{\text{moles of solute (mol)}}{\text{volume of solution (L)}}$

Molarity is moles per liter, that is, how many moles of solute (entire salt) is dissolved per liter of solution.

First, determine the saturation concentration of each of the solutions, that is, how concentrated can you get each solution before the solution is saturated. If you can't determine the concentration using the simulation "Molarity", try using the simulation "Concentration" (You will use this information again in Part 5, if your instructor requires it)



Molarity

Cobalt (II) nitrate	<u>5.640 M</u>	Potassium chromate	<u>3.350 M</u>
Cobalt chloride	<u>4.330 M</u>	Nickel (II) chloride	<u>5.210 M</u>
Potassium dichromate	<u>0.510 M</u>	Copper sulfate	<u>1.380 M</u>
Gold (III) chloride	<u>2.30 M</u>	Potassium permanganate	<u>0.480 M</u>

Concentration And Molarity Phet Chemistry Lab Answers

ML Morrison



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Table of Contents Concentration And Molarity Phet Chemistry Lab Answers

1. Understanding the eBook Concentration And Molarity Phet Chemistry Lab Answers
 - The Rise of Digital Reading Concentration And Molarity Phet Chemistry Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Concentration And Molarity Phet Chemistry Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Concentration And Molarity Phet Chemistry Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Concentration And Molarity Phet Chemistry Lab Answers
 - Personalized Recommendations
 - Concentration And Molarity Phet Chemistry Lab Answers User Reviews and Ratings
 - Concentration And Molarity Phet Chemistry Lab Answers and Bestseller Lists
5. Accessing Concentration And Molarity Phet Chemistry Lab Answers Free and Paid eBooks

- Concentration And Molarity Phet Chemistry Lab Answers Public Domain eBooks
 - Concentration And Molarity Phet Chemistry Lab Answers eBook Subscription Services
 - Concentration And Molarity Phet Chemistry Lab Answers Budget-Friendly Options
6. Navigating Concentration And Molarity Phet Chemistry Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Concentration And Molarity Phet Chemistry Lab Answers Compatibility with Devices
 - Concentration And Molarity Phet Chemistry Lab Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concentration And Molarity Phet Chemistry Lab Answers
 - Highlighting and Note-Taking Concentration And Molarity Phet Chemistry Lab Answers
 - Interactive Elements Concentration And Molarity Phet Chemistry Lab Answers
 8. Staying Engaged with Concentration And Molarity Phet Chemistry Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concentration And Molarity Phet Chemistry Lab Answers
 9. Balancing eBooks and Physical Books Concentration And Molarity Phet Chemistry Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concentration And Molarity Phet Chemistry Lab Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Concentration And Molarity Phet Chemistry Lab Answers
 - Setting Reading Goals Concentration And Molarity Phet Chemistry Lab Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Concentration And Molarity Phet Chemistry Lab Answers
 - Fact-Checking eBook Content of Concentration And Molarity Phet Chemistry Lab Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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