Molarity Calculations

Calculate the molarities of the following solutions:

 2.3 moles of sodium chloride in 0.45 liters of solution. Maci

$$M = \frac{2.3 \text{ moles of sodium chloride in 0.45 liters of solution}}{0.45 \text{ moles}} = 5.1117 \frac{mol}{1} = 5.11 \text{ M}$$

1.2 moles of calcium carbonate in 1.22 liters of solution. Cacon

0.09 moles of sodium sulfate in 12 mL of solution. NINSOU

LiF 0.75 moles of lithium fluoride in 65 mL of solution.

Ma(CH, COO),5) 0.8 moles of magnesium acetate in 5 liters of solution.

(a(NO2), B) 120 grams of calcium nitrite in 240 mL of solution.

240mL=0.240L
$$n = 1209 \left(\frac{1 \text{mod}}{132.0499} \right)$$
 $M = \frac{0.91 \text{ mod}}{0.240 \text{ L}} = \boxed{3.79 \text{ M}}$

MW_{C(NO₂)} = 13.2.099/met $N = 0.91 \text{ mod}$ $M = \frac{0.91 \text{ mod}}{0.240 \text{ L}} = \boxed{3.79 \text{ M}}$

98 grams of sodium hydroxide in 2.2 liters of solution. NAOH

1.2 grams of hydrochloric acid in 25 mL of solution.

45 grams of ammonia in 0.75 L of solution.

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Chemical Education Material Study

<u>Student Handbook</u> Southwestern,2005 The Student Handbook is designed to provide students with ready access to information with problem solving techniques and study skill guides that enable them to utilize the information in the most efficient manner Amazon com *Chemistry, an Experimental Science* Chemical Education Material Study,1963

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