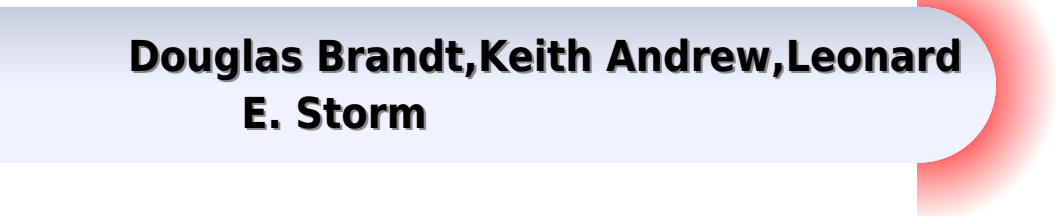


Calorimetry Practice Problems

1. How much energy is needed to change the temperature of 50.0 g of water by 15.0°C?
2. How many grams of water can be heated from 20.0 °C to 75°C using 12500.0 Joules?
3. What is the final temperature after 840 Joules is absorbed by 10.0g of water at 25.0°C?
4. The heat capacity of aluminum is 0.900 J/g°C.
 - a. How much energy is needed to raise the temperature of a 8.50×10^2 g block of aluminum from 22.8°C to 94.6°C?
5. A piece of metal weighing 59.047 g was heated to 100.0 °C and then put it into 100.0 mL of water (initially at 23.7 °C). The metal and water were allowed to come to an equilibrium temperature, determined to be 27.8 °C. Assuming no heat lost to the environment, calculate the specific heat of the metal. (Hint: First calculate the heat absorbed by the water then use this value for "Q" to determine the specific heat of the metal in a second calculation)
6. In a coffee-cup calorimeter, 100.0 g of H₂O and 100.0 mL of HCl are mixed. The HCl had an initial temperature of 44.6 °C and the water was originally at 24.6 °C. After the reaction, the temperature of both substances is 31.3 °C.
 - a. Was the reaction exothermic or endothermic? Explain.
 - b. Calculate how much heat the water lost or gained.

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**Douglas Brandt, Keith Andrew, Leonard
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