

## DEFINING THE ATOM

All numerical answers have been rounded to the correct number of significant figures.

1. How was an electric field used to determine the charge of a cathode ray?

The cathode ray was attracted to the positive end of the magnet implying the cathode ray is negative.

2. Describe a cathode-ray tube and how it operates. The cathode ray tube operates by heating a cathode to release electrons, which are then accelerated towards an anode. The anode attracts the electrons, which are then directed towards a phosphor coated screen. When the electrons hit the screen they cause it to glow, creating an image.

3. What caused the deflection of the alpha particles in Rutherford's gold foil experiment? What caused the deflection of the alpha particles in Rutherford's gold foil experiment is that The deflection of the alpha particles in Rutherford's gold foil experiment was caused by the nucleus.

4. Explain what keeps the electrons confined in the space surrounding the nucleus. The electromagnetic force is due to the attraction between the negatively charged electrons and the positively charged protons, keeps the electrons confined in the space surrounding the nucleus.

5. What are the strengths and weaknesses of Rutherford's nuclear model of the atom? The strengths are explained by the results of the gold foil experiment and why an atom is electrically neutral. The weaknesses could not account for the total mass of an atom or the arrangement of the electrons.

6. Describe the structure of a typical atom according to modern atomic theory. Identify where each subatomic particle is located. The smallest part of an element that retains the properties of the element is called an atom. It contains both protons and electrons. They are both subatomic particles. The electrons run outside of the atom, While the neutrons run inside of the atom.

7. Compare the relative charge and mass of each of the subatomic particles by filling out the table below.

| Particle | Relative Charge | Relative Mass |
|----------|-----------------|---------------|
| Electron | -1              | 1/1840        |
| Proton   | +1              | 1             |
| Neutron  | 0               | 1             |

# Defining The Atom Study Guide Answers

**R Pring**



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