

## Moles & Stoichiometry Cheat Sheet

|                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                    |                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                   |                                                             |
|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------|
| <b>Calculating Molar Mass</b><br>1. Write out formula of compound<br>2. Determine number of atoms of each element present<br>3. multiply number of atoms of the element by the atomic mass of the element<br>4. add up the values just determined for all of the elements present in the compound                                                                                                                                                                                                                                                                                                                                  | <b>Calculating Percent Composition</b><br>1. Calculate the molar mass of the compound<br>2. divide the mass contribution of each element by the total molar mass<br>3. multiply by 100 to make a percent<br><br><i>Note: total of percents should equal 100%!</i>                                                                                                                                                                                                                                 |                                                             |
| <b>Mass (grams)</b><br>x by molar mass<br>-----<br>-----<br>+ by molar mass                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                        | <b>Mole</b><br>x by Avogadro's #<br>-----<br>-----<br>+ by Avogadro's #                                                                                                                                                                                                                                                                                                                                                                                                                           | <b>Particles</b><br>(atoms,<br>molecules,<br>formula units) |
| <b>Calculating Empirical Formula</b><br>1. Change % of each element to grams of each element (if no sample amount is given assume 100 g)<br>2. convert grams of each element to moles of each element by dividing the grams by the molar (atomic) mass of that element<br>3. divide all molar values by the smallest molar value (to establish lowest whole number ratio)<br>4. if all values are not whole numbers, multiply all by a whole number factor to make the values whole (ex. 1.5 and 1 should be multiplied by 2 to get rid of the .5 = 3 and 2)<br>5. plug values in with the element they represent                  | <b>Calculating Molecular Formula</b><br>1. calculate the empirical formula first<br>2. calculate the molar mass of the empirical formula<br>3. take the molar mass of the molecular formula (given in the problem) and divide it by the molar mass of the empirical formula. This will equal the number of times the e.f. is in the m.f.<br>4. distribute the number attained through the empirical formula<br><br><i>Example: <math>2(\text{C}_2\text{H}_5) = \text{C}_4\text{H}_{10}</math></i> |                                                             |
| <b>Calculating the Formula of a Hydrate</b><br>1. Determine the mass of water in the hydrate by subtracting the mass of the anhydrous from the mass of the hydrate (if necessary) -OR- convert %'s into grams by assuming a 100g sample<br>2. convert grams of each into moles of each by dividing by molar mass<br>3. divide the moles of water by the moles of the anhydrous to determine the number of moles of water in the formula<br>4. write the formula of the anhydrous followed by a dot then the # of moles of water and "H <sub>2</sub> O"<br><br><i>Example: <math>\text{CuSO}_4 \cdot 5\text{H}_2\text{O}</math></i> | <b>Naming a Hydrate</b><br>1. write the name of the ionic compound<br>2. use a prefix for the number of water molecules (see below)<br>3. write "hydrate" after the prefix<br><br>1 = mono, 2 = di, 3 = tri, 4 = tetra, 5 = penta,<br>6 = hexa, 7 = hepta, 8 = octa, 9 = nona, 10 = deca<br><br><i>Example: Copper (II) sulfate pentahydrate</i>                                                                                                                                                  |                                                             |

# Cheat Sheet For Stoichiometry

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