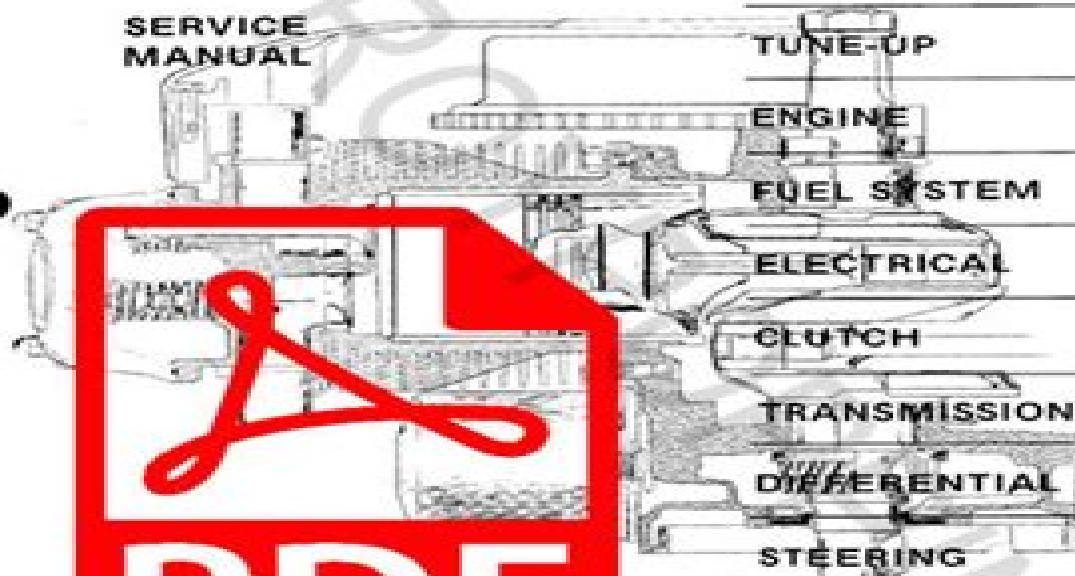


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mar 11 2021 consider these slightly soluble salts pbs with $K_{sp} = 8.4 \times 10^{-28}$ $PbSO_4$ with $K_{sp} = 1.8 \times 10^{-8}$ $PbIO_3$ with $K_{sp} = 2.6 \times 10^{-13}$ which is the most soluble calculate the solubility in moles/l for $PbSO_4$ how many grams of $PbSO_4$ dissolve in 1 l of solution

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using the graph you finished on worksheet 2 write an equation for the line depending on student ability you may provide formula for the class day 3 and 4 solubility lab overview in this lab the students will compare the dissolving rate of salt and sucrose in water at room temperature they will do this by collecting data on the time needed