

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

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Graham Booth



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Table of Contents Chemistry Study Guide Completion

1. Understanding the eBook Chemistry Study Guide Completion
 - The Rise of Digital Reading Chemistry Study Guide Completion
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Completion
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Completion
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Completion
 - Personalized Recommendations
 - Chemistry Study Guide Completion User Reviews and Ratings

- Chemistry Study Guide Completion and Bestseller Lists
- 5. Accessing Chemistry Study Guide Completion Free and Paid eBooks
 - Chemistry Study Guide Completion Public Domain eBooks
 - Chemistry Study Guide Completion eBook Subscription Services
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 - Interactive Elements Chemistry Study Guide Completion
- 8. Staying Engaged with Chemistry Study Guide Completion
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Completion
- 9. Balancing eBooks and Physical Books Chemistry Study Guide Completion
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Completion
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Study Guide Completion
 - Setting Reading Goals Chemistry Study Guide Completion
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Study Guide Completion
 - Fact-Checking eBook Content of Chemistry Study Guide Completion
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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