

chemistry basics UNIT 1

created on:
9.25.21

LAB EQUIPMENT



beaker



erlenmeyer flask



test tube



graduated cylinder



crucible



laboratory burner



thermo-meter



pipette

DENSITY

$$d = \frac{m}{v}$$

density = mass / volume
↳ g/mL, g/cm³

- cross multiply if density is given
- mL = cm³ for volume

PARTICLE DIAGRAMS



Solids



liquids



gases

most dense → least dense

Celsius °C + Kelvin
+ 273
to make
Celsius °C
Kelvin

MEASURING



= 1.10 cm

this ruler measures to the hundredths place.

↳ always round 1 more place than shown on the measuring tool

- ALWAYS view measurements from eye level
- read from the meniscus (graduated cylinder)



CONVERSIONS

- use the conversions given to find the new conversions

$$\frac{1000 \text{ mL}}{1} \times \frac{1 \text{ L}}{1000 \text{ mL}} = \frac{1000}{1000} \quad 1000 \text{ mL} = 1 \text{ liter}$$

K H D B d c m — μ — n — p
kilo hecto deca base deci centi milli nano pico
larger → smaller

* if you're converting left to right, the conversion will be smaller than the original #

significant figures

RULES:

- 0.0089040
- 340
- 340.
- 37890.00
- 0.06
- 1000

- all 0's before don't count, all after natural #'s & the decimal do.
- 0's before the tenths place will only count if there's a decimal after it.
- The 0 here counts because there is a decimal after it.
- If there are natural numbers before the decimal, anything after it counts.
- If there are no natural numbers before the decimal, sig figs start at the first natural # after the decimal.
- no decimal point after the 0's so only natural #'s count.

SCIENTIFIC NOTATION

- negative exponent = move decimal ←
↳ number will be smaller out of notation
- positive exponent = move decimal →
↳ number will be bigger out of notation

PERCENT ERROR

$$\% \text{ error} = \frac{|\text{accepted value} - \text{experimental}|}{\text{accepted value}} \times 100$$

* always include units in answers!

CHEMICAL CHANGE/PROPERTY

- change in atom formation
- underg. bond breaking / making to form something new
- anything having to do with reactions / reactivity



ex. flammability, Na is explosive in water

PHYSICAL PROPERTIES

- quality of matter that can be observed or measured without changing the bonds.
- no reactions, only things you can observe or see, or data given.



ex. the density of water is 1 g/mL, H₂O is clear.

Chemistry Of Life Note Taking Guide

Texas. State Department of Education

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products and policies just as quality and safety concerns are now addressed throughout product design and development

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Table of Contents Chemistry Of Life Note Taking Guide

1. Understanding the eBook Chemistry Of Life Note Taking Guide
 - The Rise of Digital Reading Chemistry Of Life Note Taking Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Of Life Note Taking Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Of Life Note Taking Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Of Life Note Taking Guide
 - Personalized Recommendations
 - Chemistry Of Life Note Taking Guide User Reviews and Ratings
 - Chemistry Of Life Note Taking Guide and Bestseller Lists
5. Accessing Chemistry Of Life Note Taking Guide Free and Paid eBooks

- Chemistry Of Life Note Taking Guide Public Domain eBooks
- Chemistry Of Life Note Taking Guide eBook Subscription Services
- Chemistry Of Life Note Taking Guide Budget-Friendly Options
- 6. Navigating Chemistry Of Life Note Taking Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Of Life Note Taking Guide Compatibility with Devices
 - Chemistry Of Life Note Taking Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Of Life Note Taking Guide
 - Highlighting and Note-Taking Chemistry Of Life Note Taking Guide
 - Interactive Elements Chemistry Of Life Note Taking Guide
- 8. Staying Engaged with Chemistry Of Life Note Taking Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Of Life Note Taking Guide
- 9. Balancing eBooks and Physical Books Chemistry Of Life Note Taking Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Of Life Note Taking Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Of Life Note Taking Guide
 - Setting Reading Goals Chemistry Of Life Note Taking Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Of Life Note Taking Guide
 - Fact-Checking eBook Content of Chemistry Of Life Note Taking Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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