

Chemistry Review for Final

- Classify each of the following as homogeneous or heterogeneous mixtures:
a) blood b) chocolate chip cookies c) dissolved "koolaid" d) homogenized milk e) bronze
- Classify the following changes as physical or chemical:
a) food spoils b) an icicle melts c) a nail rusts
d) oil is pumped out of a well e) an egg is fried f) salt dissolves in water
g) a window is broken
- A compound is analyzed and is determined to be made up of 5.88% hydrogen, and 94.12% oxygen by mass. What is its empirical formula?
- Given the following percent compositions and molecular weights, determine the molecular formula for each compound.
a) 81.7% carbon, 18.3% hydrogen, molecular mass = 44.1 g/mol
b) 37.8% carbon, 6.4% hydrogen, 55.8% chlorine, molecular mass = 127.0 g/mol
- Describe the different types of intermolecular forces.
- Predict the monatomic ions formed by the following elements:
a) hydrogen b) magnesium c) sulfur d) iodine e) aluminum
- Balance the following equations. If they need no coefficients, write "balanced."
a) $\text{CF}_4(\text{l}) \rightarrow \text{C}(\text{s}) + \text{F}_2(\text{g})$
b) $\text{H}_2\text{SO}_4(\text{aq}) + \text{KOH}(\text{aq}) \rightarrow \text{KH}_2\text{SO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l})$
c) $\text{ZnCl}_2(\text{aq}) + \text{H}_2(\text{g}) \rightarrow \text{Zn}(\text{s}) + \text{HCl}(\text{aq})$
d) $\text{SO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{SO}_4(\text{aq})$
e) $\text{Li}(\text{s}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{LiOH}(\text{aq}) + \text{H}_2(\text{g})$
f) $\text{Cu}_2\text{O}(\text{aq}) + \text{Cu}_2\text{S}(\text{aq}) \rightarrow \text{Cu}(\text{s}) + \text{SO}_2(\text{g})$
g) $\text{Na}_2\text{SO}_4(\text{aq}) + \text{BaCl}_2(\text{aq}) \rightarrow \text{BaSO}_4(\text{s}) + \text{NaCl}(\text{aq})$
h) $\text{CH}_3\text{OH}(\text{l}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
- Classify each of the above according to the 5 types of reactions taken in class. If no classification fits, write "other."
- Write the formula for each material correctly and then balance the equation. For each reaction, tell what type of reaction it is. For some reactions, you will need to determine the products.
 - sulfur trioxide and water combine to make sulfuric acid.
 - lead (II) nitrate and sodium iodide react to make lead iodide and sodium nitrate.
 - calcium fluoride and sulfuric acid make calcium sulfate and hydrogen fluoride (Hydrofluoric acid)
 - calcium carbonate will come apart when you heat it to leave calcium oxide and carbon dioxide.
 - propane burns (with oxygen)
 - aluminum sulfate and calcium hydroxide become aluminum hydroxide and calcium sulfate.
 - copper metal and silver nitrate react to form silver metal and copper (II) nitrate.
 - sodium metal and chlorine react to make sodium chloride.
- Describe the difference between ionic and molecular compounds. You may answer in point form.
- Classify each of the following as ionic or molecular, and give the correct chemical name for each:
 - NaI
 - FeO
 - K_2S
 - GaBr_3
 - S_4N_2
 - $\text{Zn}(\text{CH}_3\text{COO})_2$
 - SnF_4
 - $\text{LiCl} \cdot 4\text{H}_2\text{O}$
 - P_2O_5
- Give the correct chemical formula for each of the following:
 - hydrochloric acid
 - sodium hexafluoride

Chemistry Final Review

Scott C. Dulebohn



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her interactions with many physicians she met working at the hospital mostly expatriates She also describes her vacation travels during that time in and around Ethiopia and also to Kenya Tanzania and Uganda She visited several game parks and climbed Mt Kilimanjaro Goodwin also writes about her interactions fellow college teachers Peace Corps volunteers and Ethiopians She describes several instances of anti Peace Corps agitation in Ethiopia especially its effect on the local secondary school

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