

Equilibria Test SCH4U Name:

15 K 10 T 11 C 10 A Assessment of Learning (Product)

Multiple Choice: Circle the most correct answer. (15K)

1. This symbol () indicates that _____

- A) heat must be applied B) an incomplete combustion reaction has occurred
C) a gas is formed by the reaction D) the reaction is reversible

2. Which one of the following will change the value of an equilibrium constant?

- A) changing temperature
B) adding other substances that do not react with any of the species involved in the equilibrium
C) varying the initial concentrations of reactants
D) varying the initial concentrations of products

3. The K_p for the reaction below is 1.49×10^2 at 100.0°C : $\text{CO(g)} + \text{Cl}_2\text{(g)} \rightleftharpoons \text{COCl}_2\text{(g)}$

In an equilibrium mixture of the three gases, $P_{\text{CO}} = P_{\text{Cl}_2} = 8.60 \times 10^{-4} \text{ atm}$. The partial pressure of the product, phosgene (COCl_2), is _____ atm.

- A) 1.1×10^2 B) 2.01×10^{14} C) 4.96×10^{-15} D) 1.28×10^5 E) 172×10^{11}

4. Dinitrogen tetroxide partially decomposes according to the following equilibrium:

$\text{N}_2\text{O}_4\text{(g)} \rightleftharpoons 2\text{NO}_2\text{(g)}$ A 1.000-L flask is charged with $3.00 \times 10^{-2} \text{ mol}$ of N_2O_4 . At equilibrium, $2.36 \times 10^{-2} \text{ mol}$ of N_2O_4 remains. K_c for this reaction is _____

- A) 0.723 B) 0.391 C) 0.212 D) 6.94×10^{-3} E) 1.92×10^{-3}

5. At 400 K, the equilibrium constant for the reaction $\text{Cl}_{2(g)} + \text{Br}_{2(g)} \rightleftharpoons 2\text{BrCl}_{(g)}$

is $K_p = 7.0$. A closed vessel at 400 K is charged with 1.00 atm of $\text{Br}_2\text{(g)}$, 1.00 atm of $\text{Cl}_2\text{(g)}$, and 2.00 atm of BrCl(g) . Use Q to determine which of the statements below is true

A) The equilibrium partial pressures of Br_2 , Cl_2 , and BrCl will be the same as the initial values.

B) The equilibrium partial pressure of Br_2 will be greater than 1.00 atm.

C) At equilibrium, the total pressure in the vessel will be less than the initial total pressure

D) The equilibrium partial pressure of BrCl(g) will be greater than 2.00 atm.

E) The reaction will go to completion since there are equal amounts of Br_2 and Cl_2 .

Chemistry Equilibrium Test Answers

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