

EQUILIBRIUM CHEAT SHEET

Chemical equilibrium, dynamic equilibrium, where the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium reactions are represented by this symbol: $\rightleftharpoons$

What is Equilibrium?

Equilibrium is a constant concentration of reactants

When a reversible reaction takes place in a closed system, the forward and reverse reactions occur at the same rate. The system is called dynamic equilibrium when the concentrations of reactants and products do not change.

It is interesting to think of equilibrium as a constant state. The forward and reverse reactions are still occurring, but the concentrations of reactants and products are constant. The system is in a state of dynamic equilibrium.

Equilibrium Constant Formed

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Equilibrium Constant, K_c

The equilibrium constant, K_c , is a ratio of the concentrations of products to reactants, raised to the power of their stoichiometric coefficients. It is a constant for a given reaction at a given temperature.

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Equilibrium Constant Equation

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

The equilibrium constant, K_c , is a ratio of the concentrations of products to reactants, raised to the power of their stoichiometric coefficients. It is a constant for a given reaction at a given temperature.

Reaction Quotient, Q

The reaction quotient, Q , is a ratio of the concentrations of products to reactants, raised to the power of their stoichiometric coefficients. It is a constant for a given reaction at a given temperature. It is used to determine if a reaction is at equilibrium. If $Q = K_c$, the reaction is at equilibrium. If $Q < K_c$, the reaction is not at equilibrium and will proceed in the forward direction. If $Q > K_c$, the reaction is not at equilibrium and will proceed in the reverse direction.

$$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Equilibrium Constant

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Example

$$2 H_2 + O_2 \rightleftharpoons 2 H_2O$$

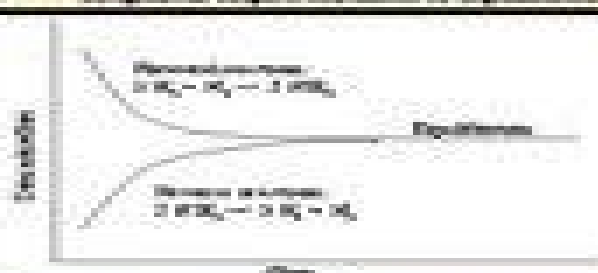
Forward Reaction

$$2 H_2 + O_2 \rightarrow 2 H_2O$$

Reverse Reaction

$$2 H_2O \rightarrow 2 H_2 + O_2$$

Graphical Representation of Equilibrium



Equilibrium Constant Relationship

Equilibrium Constant Value	Meaning
$K_c > 1$	Concentration of products is greater than concentration of reactants
$K_c < 1$	Concentration of reactants is greater than concentration of products
$K_c = 1$	Concentration of reactants is equal to concentration of products

Reaction Quotient vs. Equilibrium Constant

Equilibrium Constant vs. Reaction Quotient	Meaning
$Q = K_c$	Reaction is at equilibrium. $2 H_2 + O_2 \rightleftharpoons 2 H_2O$
$Q < K_c$	Reaction is not at equilibrium. Forward reaction is favored. $2 H_2 + O_2 \rightarrow 2 H_2O$
$Q > K_c$	Reaction is not at equilibrium. Reverse reaction is favored. $2 H_2O \rightarrow 2 H_2 + O_2$

Le Chatelier's Principle: The effect of changing conditions on equilibrium

Condition	Effect of Change	Example
Concentration	- Increasing concentration of reactants shifts equilibrium to the right. - Increasing concentration of products shifts equilibrium to the left.	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing concentration of H_2 or O_2 shifts equilibrium to the right, increasing concentration of H_2O .
Pressure	Increasing pressure shifts equilibrium to the side with fewer moles of gas.	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing pressure shifts equilibrium to the right, increasing concentration of H_2O .
Temperature	Increasing temperature shifts equilibrium to the side with higher energy (endothermic reaction).	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing temperature shifts equilibrium to the left, decreasing concentration of H_2O .

Chemistry Chemical Equilibrium Study Guide Answers

Lei Huang



Chemistry Chemical Equilibrium Study Guide Answers:

The book delves into Chemistry Chemical Equilibrium Study Guide Answers. Chemistry Chemical Equilibrium Study Guide Answers is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemistry Chemical Equilibrium Study Guide Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chemistry Chemical Equilibrium Study Guide Answers
 - Chapter 2: Essential Elements of Chemistry Chemical Equilibrium Study Guide Answers
 - Chapter 3: Chemistry Chemical Equilibrium Study Guide Answers in Everyday Life
 - Chapter 4: Chemistry Chemical Equilibrium Study Guide Answers in Specific Contexts
 - Chapter 5: Conclusion
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