Test Date Name KEY Period

CP Chemistry - Final Exam Review

Show your work and round using significant figure rules for all the problems.

- 1. What is chemistry?
 - Chemistry is the study of matter and the changes it undergoes.
- 2 Write the following in scientific notation:
 - 5 x 10° a = 0.05

 5.3099×10^{7} c = 530990000

b. 5050. 5.05 x 10° d. 0.00053030 5.3030 x 10⁻⁶

- 3. Write the following in standard notation:
 - a. 1.5×10^3

- c. 4x10⁻⁵
- 0.00004

b. 3.75×10^{-2} 0.0375

- d. 9.12 x10⁷ 91200000
- 40 How many significant figures are in the following:
 - a. 0.02

1

1500

- c. 501

b. 501.0

- d = 0.020
- 5. Calculate and report in correct significant figures:
 - a. 52.568 cm + 689.45 cm + 55.5694 cm = 797.59 cm
 - b. 8426, 200 g + 254 mol = 33.2 g /mol
- You measure a Calcium cube to have 2.01 cm sides and a mass of 11.6 grams: 6.
 - a. What is the density as you calculate it?

$$y = s^3 = (2.01 \text{ cm})^3 = 8.12 \text{ cm}^3$$

 $d = \frac{m}{v} = \frac{11.6 \text{ g}}{8.12 \text{ cm}^3} = 1.43 \text{ g/cm}^3$

b. If calcium's accepted density is 1.54 g/cm³, what is your percent error?

$$\%_{error} = \frac{|1.54 \text{ g/cm}^3 - 1.43 \text{ g/cm}^3|}{1.54 \text{ g/cm}^3} \times 100 = 7.14 \%$$

- 7. Perform the following conversions:
 - 53.1 mm to m

53.1 mm x
$$\frac{1 \text{ m}}{10^3 \text{ mm}} = 0.0531 \text{ m}$$

b. 18.4 kg to dag

$$18.4 \text{ kg x} \frac{1 \text{ x } 10^{-1} \text{ dag}}{1 \text{ x } 10^{-3} \text{ kg}} = 1840 \text{ dag}$$

c. 93.6 mi to km (1 mi = 1.6 km)

93.6 mi x
$$\frac{1.6 \text{ km}}{1 \text{ mi}}$$
 = 150, km

Chemistry Spring Semester Final Study Guide

ML Morrison

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