Experiment 18-Chemical Reactions Lab Report

Name:		
vaille:	Section:	

For this lab, you need to balance each equation, and classify the reaction according to the types of reactions listed in the lab handout using C (combination), D (decomposition), Cu (combustion), S (single replacement) and DR (double replacement). Include what you observed below each reaction.

1. __CaCl₂ (aq) + __Na₃PO₄ (aq) → __Ca₃(PO₄)₂ (s) + __NaCl (aq) _____

Observations:

2. __CuSO₄ (H₂O)₅ + heat → __CuSO₄ (s) + __H₂O (g) _____ Observations:

__CuSO₄ (s) + __H₂O (l) → __CuSO₄ (H₂O)₅ (s) _____ Observations:

3. __Zn (s) + __HCl (aq) → __ZnCl₂ (aq) + __H₂ (g) _____ Observations:

 $_H_2(g) + _O_2(g) \rightarrow _H_2O(g)$ The "popping" reaction induced with the match. _____Observations:

4. __HCl(aq) + __Na₂CO₃(aq) \rightarrow __CO₂(g) + __NaCl(aq) + __H₂O(l) _____ Observations:

H₂O₂(aq) → __O₂(g) + __H₂O(l) (the MnO₂ is a catalyst, it does not react) _____
 Observations:

6. __HCl (aq) + __NaOH (aq) → __NaCl (aq) + __H₂O (l) _____

7. __CaCl₂ (aq) + __Na₂CO₃ (aq) → __NaCl (aq) + __CaCO₃ (s) _____ Observations:

8. __Zn (s) + __CuSO₄ (aq) → __ZnSO₄ (aq) + __Cu (s) _____

9. __Cu (s) + __ZnSO₄ (aq) → __CuSO₄ (aq) + __Zn (s) _____ Observations:

10. __CaCl₂ (aq) + __NaNO₃ (aq) → __Ca(NO₃)₂ (aq) + __NaCl (aq) _____

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