# Acids and Bases



Acid				Base		
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base
Strongest	Large	Perchloric acid	HCIO.	CIO,-	Perchlorate ion	Weakest
	3.2×10*	Hydroiodic acid		1	Todide	
	1.0×10*	Hydrobromic acid	HBr	Bre	Bromide	T 72
	$1.3 \times 10^{4}$	Hydrochloric acid	HCI	CI-	Chloride	
	1.0×100	Sulfuric acid	H <sub>i</sub> SO <sub>4</sub>	HSO,-	Hydrogen sulfate ion	
	2.4×10 <sup>1</sup>	Mitric acid	HNO <sub>2</sub>	NO <sub>y</sub> -	Nitrate ion	
		Hydronium ion	H <sub>2</sub> O+	нуо	Water	
	5.4×10 <sup>-3</sup>	Oxelic acid	но,с,о,н	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> -	Hydrogen exalate ion	
	1.3×10-	Sulfurous acid	H <sub>a</sub> SO <sub>a</sub>	HSO,-	Hydrogen sulfite ion	
	1.0×104	Hydrogen sulfate ion	HSO: -	50.3	Sulfate ion	-
	$7.1 \times 10^{-3}$	Phosphoric acid	н,РО.	H <sub>2</sub> PO <sub>2</sub> -	Dihydrogen phosphate ion	
	7.2×10 <sup>-4</sup>	Nitrous acid	HNO <sub>a</sub>	NO <sub>2</sub> -	Nitrite ion	
	6.6×10*	Hydrofluoric acid	HF	F-	Fluoride ion	
	1.8×10-	Methanoic acid	нооун	HCO <sub>2</sub> -	Methanoate ion	
	6.3×10 <sup>-6</sup>	Benzoic acid	С"Нускон	C,H,000-	Benzoate ion	
Medium	5.4×10°	Hydrogen oxalate ion	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> -	O <sub>2</sub> C <sub>2</sub> O <sub>2</sub> 2-	Oxalate ion	Medium
	1.8×10 <sup>-5</sup>	Ethanoic acid	СН,СООН	CH/COO	Ethanoate (acetate) ion	
	$46.48 \pm 1.0^{-7}$	Carbonic acid	н, со,	HCO <sub>3</sub> -	Hydrogen carbonate ion	
	$1.1 \times 10^{-r}$	Hydrosulfuric acid	H <sub>1</sub> S	HS-	Hydrogen sulfide ion	
	6.3×10°	Dihydrogen phosphate ion	H <sub>2</sub> PO <sub>4</sub> +	HPO <sub>x</sub> 2-	Hydrogen phosphate ion	
	6.2×10*	Hydrogen suifite ion	HS-	50,1	Sulfite ion	
	$2.9 \times 10^{-6}$	Hypochlorous acid	HCIO	CIO-	Hypochlorite ion	
	6.2×10-4	Hydrocyanic acid	HCN	CN-	Cyanide ion	
	5.8×10 <sup>++</sup>	Ammonium ion	NHL +	NH <sub>3</sub>	Ammonia	
	5.8×10 **	Boric acid	H,80,	H <sub>2</sub> BO <sup>2</sup>	Dihydrogen borate	
	$4.7 \times 10^{-11}$	Hydrogen carbonate ion	HCO <sub>2</sub> -	CO2F	Carbonate ion	
	4.2×10 <sup>-13</sup>	Hydrogen phosphate ion	HPO, P	PO. P	Phosphate ion	
	1.8×10-45	Dihydrogen borate ion	н,воз-	нво, -	Hydrogen borate ion	
	$1.3\!\times\! 10^{-13}$	Hydrogen sulfide ion	HS-	S.1-	Sulfide ion	1
	1.6×10-0*	Hydrogen borate ion	нво, -	BO, *	Borate ion	
Weekest		water	H <sub>2</sub> O	OH-	Hydroxide	Stronges



# **Chemistry Acids And Bases Names**

**JE Gale** 

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