

1.9: Solutions

Key Solution Terms

- Solute:** Substance dissolved. The substance formation produces the breaking and forming of intermolecular forces. $\text{Solute} + \text{Solvent} = \text{Solution}$.
- Solvent:** Substance being dissolved (or smaller amount).
- Solvent:** Substance doing the dissolving (or larger amount).
- Solute vs Solvent Measurement:** Creative units for solute vs solvent that must make that ratio in the past.
- Unsaturated Solution:** The solution can hold more solute particles.
- Saturated Solution:** The solution is holding as many solute particles as it can.
- Supersaturated Solution:** The solution contains more solute particles than it should be able to at that temperature. It can happen with solute supersaturation by cooling the temperature, which allows more solute to dissolve, and then heating the temperature slowly kept down.
- Concentration:** ratio of solute to solvent or solution.
- Concentration:** high ratio of solute to solvent.
- Solubility:** Capacity showing the solubility of a solute in a solvent, temperature.
- Solubility:** rate of solute to solvent.
- Solubility:** ability to create a more stable solution.
- Electrolyte:** compound that dissociates into ions when dissolved in water. Electrolyte is conductive electricity.
- Electrolyte:** solution with solute particles large enough to conduct and transfer effect.
- Crystal effect:** The light is visible and scattered as it travels through the solid.

Formation of Solutions

Steps for solution formation:

1. "Expand the solvent" - break intermolecular forces within the solvent.
2. "Expand the solute" - break intermolecular forces within the solute.
3. Form new intermolecular forces between solute & solvent.
4. Steps 1 & 2 require energy. Step 3 releases energy.
5. If the energy put in is more greater than the energy released the system will not form.

Factors Affecting Solubility

"Like Dissolves Like"

- Solubility:** with "like" bond properties will form similar intermolecular forces with each other as they are like themselves. Therefore, the energy released will be similar to the energy put in.

Pressure:

- Gas:** dissolve more when there is a higher pressure of the gas above the solution.

Temperature:

- Gas:** dissolve more when the temperature of the solution is lower.
- Solid:** solute dissolve more when the temperature of the solution is higher.

How to Use This Cheat Sheet: Focus on the blue shading for the facts. The yellow shading is possible facts that are not as solid as the blue shading of facts. Focus on your learning objectives.

Concentrations of Solutions

By the mass: is mass of solute divided by mass of solution.
 The mass will most likely.

By the volume: is volume of solute divided by volume of solution.
 The volume will most likely.

By the molarity: is moles of solute divided by volume of solution.
 The volume will be in L.

Molarity (M): moles of solute divided by volume of solution.

Molarity (M): moles of solute divided by volume of solution.

Molarity vs Molar Mass: Molarity (moles per liter solution) and Molar Mass (grams per mole solute) = "Moles per gram per liter of solute".

Calculations with Concentrations

Solution equation: $(M_1 V_1) = (M_2 V_2)$
 M_1 = original molarity M_2 = original volume
 V_1 = new molarity V_2 = new volume
 (moles solute must match)

Solute Dissolving

How dissolving for form the equation: solvent must add more of a solution.

Example: How many grams of 1.7M NaCl are needed to make 2.5 L of NaCl?

$$1.7 \text{ M NaCl} \times 2.5 \text{ L} = 4.25 \text{ moles NaCl} \times 58.44 \text{ g/mol} = 248.3 \text{ g NaCl}$$

Electrolyte Solutions

The solute can be electrolyte.

- Electrolyte:** solute that dissociates into ions.
- Electrolyte:** solute that dissociates into ions.

Example:

$\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$
 $\text{CaCl}_2 \rightarrow \text{Ca}^{2+} + 2 \text{Cl}^-$

Colligative

Colligative: solution with large enough particles to conduct light.

Crystal effect:

For a solution: 

For a solid: 

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