

Gas Laws and Theory SCH3U

Name:

$\frac{\text{K}}{8}$ $\frac{\text{T}}{10}$ $\frac{\text{C}}{4}$ $\frac{\text{A}}{3}$ Assessment of Learning (Product)

$$R = \frac{8.31 \text{ kPa} \times \text{L}}{\text{Mol} \times \text{K}}$$

$$1 \text{ atm} = 101.325 \text{ kPa} = 760 \text{ torr} = 760 \text{ mmHg}$$

$$\text{SATP} = 25^\circ\text{C and } 100 \text{ kPa}$$

$$\text{STP} = 0^\circ\text{C and } 101.3 \text{ kPa}$$

Multiple Choice. Circle the most correct answer. (8K)

1. The volume of a confined gas can be lowered by increasing the pressure acting upon the gas (assuming constant temperature). The difference in volume is best explained by which property of gas molecules?

- a. Gas molecules have high vibrational motion.
- b. Gas molecules move in straight line motion.
- c. Gas molecules are considered dimensionless points.
- d. Gas molecules have elastic collisions with one another.
- e. Gas particles are very far apart at low pressure.

2. The particles in a liquid are usually

- A) closer together and lower in energy than those in a solid
- B) farther apart and higher in energy than those in a gas.
- C) closer together and lower in energy than those in a gas.
- D) farther apart and lower in energy than those in a solid

3. Pressure is the force per unit

- A) volume
- B) surface area
- C) length
- D) depth

4. Why does the pressure inside a container of gas increase if more gas is added to the container?

- A) There is an increase in the number of collisions between particles and the walls of the container.
- B) There is an increase in the temperature of the gas.
- C) There is a decrease in the volume of the gas.
- D) There is an increase in the force of the collisions between the particles and the walls of the container.

5. What happens to the temperature of a gas when it is compressed?

- A) The temperature increases.
- B) The temperature does not change
- C) The temperature decreases.
- D) The temperature becomes unpredictable

6. The volume of a gas is reduced from 4 L to 0.5 L while the temperature is held constant. How does the gas pressure change?

- A) It increases by a factor of four.
- C) It increases by a factor of eight.
- B) It decreases by a factor of eight.
- D) It increases by a factor of two.

7. Convert the pressure 0.840 atm to mm Hg.

- A) 365 mm Hg
- B) 437 mm Hg
- C) 638 mm Hg
- D) 780 mm Hg

Chemistry Gas Laws Quiz Test Answer

David Baud



Chemistry Gas Laws Quiz Test Answer:

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Table of Contents Chemistry Gas Laws Quiz Test Answer

1. Understanding the eBook Chemistry Gas Laws Quiz Test Answer
 - The Rise of Digital Reading Chemistry Gas Laws Quiz Test Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Gas Laws Quiz Test Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Gas Laws Quiz Test Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Gas Laws Quiz Test Answer
 - Personalized Recommendations
 - Chemistry Gas Laws Quiz Test Answer User Reviews and Ratings
 - Chemistry Gas Laws Quiz Test Answer and Bestseller Lists
5. Accessing Chemistry Gas Laws Quiz Test Answer Free and Paid eBooks
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 - Chemistry Gas Laws Quiz Test Answer Compatibility with Devices
 - Chemistry Gas Laws Quiz Test Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Gas Laws Quiz Test Answer
 - Highlighting and Note-Taking Chemistry Gas Laws Quiz Test Answer
 - Interactive Elements Chemistry Gas Laws Quiz Test Answer
8. Staying Engaged with Chemistry Gas Laws Quiz Test Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Gas Laws Quiz Test Answer
9. Balancing eBooks and Physical Books Chemistry Gas Laws Quiz Test Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Gas Laws Quiz Test Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Gas Laws Quiz Test Answer
 - Setting Reading Goals Chemistry Gas Laws Quiz Test Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Gas Laws Quiz Test Answer
 - Fact-Checking eBook Content of Chemistry Gas Laws Quiz Test Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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