

### Balancing Equations and Simple Stoichiometry-KEY

Balance the following equations:

- 1)  $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
- 2)  $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$
- 3)  $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$
- 4)  $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$
- 5)  $3 \text{ SnO} + 2 \text{ NF}_3 \rightarrow 3 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3$

Using the following equation:



- 6) How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

$$? \text{ g Na}_2\text{SO}_4 = 200 \text{ g NaOH} \rightarrow \text{mol NaOH} \rightarrow \text{mol Na}_2\text{SO}_4 \rightarrow \text{g Na}_2\text{SO}_4$$

$$200 \text{ g NaOH} \times \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} = 5 \text{ mol NaOH} \times \frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} = 2.5 \text{ mol Na}_2\text{SO}_4$$

$$2.5 \text{ mol Na}_2\text{SO}_4 \times \frac{142 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = \boxed{355 \text{ g Na}_2\text{SO}_4}$$

- 7) Using the following equation:



How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

Molar masses:  $\text{LiNO}_3 = 68.95 \text{ g/mol}$  &  $\text{Li}_2\text{SO}_4 = 109.9 \text{ g/mol}$

$$? \text{ g LiNO}_3 = 250 \text{ g Li}_2\text{SO}_4 \rightarrow \text{mol Li}_2\text{SO}_4 \rightarrow \text{mol LiNO}_3 \rightarrow \text{g LiNO}_3$$

$$250 \text{ g Li}_2\text{SO}_4 \times \frac{1 \text{ mol Li}_2\text{SO}_4}{109.9 \text{ g Li}_2\text{SO}_4} = 2.27 \text{ mol Li}_2\text{SO}_4 \times \frac{4 \text{ mol LiNO}_3}{2 \text{ mol Li}_2\text{SO}_4} = 4.54 \text{ mol LiNO}_3$$

$$4.55 \text{ mol LiNO}_3 \times \frac{68.95 \text{ g LiNO}_3}{1 \text{ mol LiNO}_3} = \boxed{314 \text{ g LiNO}_3}$$

# Chemfiesta Stoichiometry Lab Answers

**Luann Marie Decker**

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