

Density and Specific Gravity

3

A. Density of Liquids

A.1 Density of Water

1. Volume (mL)
2. Mass of beaker
3. Mass of beaker + liquid
4. Mass of liquid
5. Density

(Show calculations)

$$\begin{array}{r}
 20.0 \text{ mL} \\
 0.00 \text{ g} \\
 15.82 \text{ g} \\
 \hline
 20.420 \text{ g} \\
 D = \frac{m}{V} = \frac{20.420}{20.0} = 1.01 \text{ g/mL}
 \end{array}$$

A.2 Density of an Unknown Liquid

1. Type of liquid sample
Volume of liquid
2. Mass of beaker
3. Mass of beaker + liquid
4. Mass of liquid
5. Density

(Show calculations)

$$\begin{array}{r}
 \text{Salt water} \\
 20.1 \text{ mL} \\
 0.00 \text{ g} \\
 15.89 \text{ g} \\
 \hline
 21.24 \text{ g} \\
 \frac{21.24 \text{ g}}{20.1 \text{ mL}} = 1.05 \text{ g/mL}
 \end{array}$$

B. Specific Gravity

1. Specific gravity of water
2. Specific gravity of liquid sample
3. Specific gravity of water using a hydrometer
4. Specific gravity of liquid using a hydrometer

$$\begin{array}{r}
 1.01 \\
 1.05 \\
 0.995 \\
 0.790
 \end{array}$$

How do the *calculated* specific gravity of water and the liquid sample (1, 2) compare to the hydrometer readings (3, 4) for water and the liquid sample?

Density And Specific Gravity Laboratory Manual

Lester L. Skolil, Louis E. Smith



Density And Specific Gravity Laboratory Manual:

A Laboratory Manual of Physics Caleb Canby Balderston, 1898 **Physical Laboratory Manual** Charles Francis Adams, 1909 **A Laboratory Manual of Physics and Applied Electricity** Edward Leamington Nichols, 1894 **A Laboratory manual of physics and applied electricity v. 1**, 1894 **Laboratory Experiments to Accompany General, Organic and Biological Chemistry** Charles Anderson, David B. Macaulay, Molly M. Bloomfield, Joseph M. Bauer, 2013-02-04

This General Organic and Biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry An integrated approach is employed in which related general chemistry organic chemistry and biochemistry topics are presented in adjacent chapters This approach helps students see the strong connections that exist between these three branches of chemistry and allows instructors to discuss these interrelationships while the material is still fresh in students minds

A Laboratory Manual of Experiments in Physics Leonard Rose Ingersoll, 1925 *A Laboratory Manual of Physics and Applied Electricity: Junior course in general physics, by Ernest Merritt and Frederick J. Rogers* Edward Leamington Nichols, 1897 **Physics Laboratory Manual** William D. Henderson, 1930 *Modern Physics Laboratory Manual* Lester L. Skolil, Louis E. Smith, 1973-06 **A Laboratory Manual of Physics for Use in High Schools** Henry Crew, Robert Richardson Tatnall, 1902

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