



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM



EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the reaction $\rightleftharpoons$.

What is Equilibrium? Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the reaction $\rightleftharpoons$.		
Equilibrium Constant, K_c The equilibrium constant, K_c , is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where $[A]$, $[B]$, $[C]$, and $[D]$ are the molar concentrations of the species at equilibrium.		
Equilibrium Constant, K_p The equilibrium constant, K_p , is a ratio of the partial pressures of the products to the partial pressures of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$ where P_A , P_B , P_C , and P_D are the partial pressures of the species at equilibrium.		
Equilibrium Constant, K The equilibrium constant, K , is a general term that encompasses both K_c and K_p . It is a ratio of the concentrations or partial pressures of the products to the concentrations or partial pressures of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ or $K = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$		
Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Change	Effect on Equilibrium	Result
Concentration	Increasing the concentration of a reactant or product shifts the equilibrium in the direction that consumes that species.	Shifts in the direction of the change.
Pressure	For reactions involving gases, increasing the pressure shifts the equilibrium in the direction that produces fewer moles of gas.	Shifts in the direction of the change.
Temperature	- Exothermic reactions: increasing the temperature shifts the equilibrium in the direction that produces heat (the reactants). - Endothermic reactions: increasing the temperature shifts the equilibrium in the direction that consumes heat (the products).	Shifts in the opposite direction of the change.

Chemical Equilibrium Study Guide For Content Mastery

Steven S. Zumdahl, Paul B. Kelter



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