

Organic Chemistry

Questions and Answers

2018
Edition

ene is large. Aromaticity in naphthalene is also large. Aromaticity has come to mean a stabilization resulting from (p-orbital) overlap in a pi-type system. As the examples are large or small.

Let's consider C_{60} :

Bond alternation is often taken as a sign of non-aromatic systems. In C_{60} , bond lengths, ~1.4 and 1.45 angstroms. However, this variation is on the order of that found in polycyclic aromatic hydrocarbons, and less than that observed in non-aromatic systems. Conclusion: aromatic, but less so than benzene.

Magnetic properties are related to electron delocalization and are often used to measure aromaticity. Both experiment and calculations suggest the existence of ring currents (diamagnetic and paramagnetic) in C_{60} .

Conclusion: Although analysis is complex, analysis is consistent with C_{60} being aromatic to some degree of aromaticity.

Reactivity - Substitution reactions are not possible as no hydrogens are attached. If an anion or radical is added to C_{60} the electron(s) are not delocalized over the entire structure. However, most addition reactions are reversible suggesting that C_{60} has some stability or aromaticity associated with it.

Conclusion: Not as aromatic as benzene

Resonance energy calculations have been performed and give conflicting results. Some suggest a small stabilization. Theoretical analysis of the following isomerization reaction suggests that it only took half as much energy to break a C-C bond as in benzene. $C_{60} + 120 CH_4 \rightarrow 30 C_2H_4 + 60 C_2H_2$

Conclusion: Some aromaticity is present. by George Duckett

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Mark Steyvers



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