

1. One cause of low-level smog is the reaction of ozone, O_3 , with ethene, C_2H_4 . The smog contains methanal, $\text{HCHO}(\text{g})$.

The equation for methanal production is shown below.



The rate of the reaction was investigated, using a series of different concentrations of either $\text{C}_2\text{H}_4(\text{g})$ or $\text{O}_3(\text{g})$, by measuring the initial rate of formation of $\text{HCHO}(\text{g})$.

The results are shown below.

experiment	$[\text{O}_3(\text{g})]$ / $10^{-7} \text{ mol dm}^{-3}$	$[\text{C}_2\text{H}_4(\text{g})]$ / $10^{-8} \text{ mol dm}^{-3}$	initial rate / $10^{-12} \text{ mol dm}^{-3} \text{ s}^{-1}$
1	0.5	1.0	1.0
2	2.0	1.0	4.0
3	4.0	2.0	16.0

- (i) Analyse and interpret the results to deduce the order of reaction of each reactant and the rate equation.

Explain your reasoning.

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Camilla Rothe



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