Acids and Bases



Acid				Base		
Strength of Acid	Ka	Name	Formula	Formula	Name	Strengti of Base
trongest	Large	Perchloric acid	HCIO,	CIO,-	Perchlorate ion	Weakest
	3.2×10*	Hydrolodic acid		1-	Iodide	
	1.0×10*	Hydrobromic acid	HBr	Br-	Bromide	
	$1.3 \times 10^{\circ}$	Hydrochloric acid	HCI	CI-	Chloride	
	1.0×10°	Sulfuric acid	H,SO.	HSO	Hydrogen sulfate ion	
	2.4×10°	Nitric acid	HNO ₃	NO _y -	Nitrate ion	
		Hydronium ion	H ₁ O+	H ₂ O	Water	
	5.4×10°	Oxelic acid	HO ₁ C ₂ O ₂ H	HO ₂ C ₂ O ₂ -	Hydrogen oxalate ion	
	1.3×10-	Sulfurous acid	H ₃ SO ₃	HSO ₃ -	Hydrogen sulfite ion	
	1.0=10-	Hydrogen sulfate ion	HSO ₂ -	50, 4	Sulfate ion	
	7.4×10^{-3}	Phosphoric acid	н, Ро.	H ₂ PO ₂ -	Dihydrogen phosphate ion	
	7.2×10*	Nitrous acid	HNO,	NO ₂ -	Nitrite ion	
	6.6×10*	Hydrofluoric acid	HF	F-	Fluoride ion	
	1.8×10 ⁻⁴	Methanoic acid	HCO ₂ H	HCO ₂ -	Methanoate ion	
	6.3×10°	Benzoic acid	С"Н,СООН	C,H,000-	Benzoate ion	
Medium	5.4×10°	Hydrogen oxalate ion	HO ₂ C ₂ O ₂ -	O ₂ C ₂ O ₂ 2-	Oxalate ion	Medium
	1.8×10-4	Ethanoic acid	СН,СООН	CH,COO	Ethanoate (acetate) ion	
	4.4×10^{-7}	Carbonic acid	н, со,	HCO ₂ -	Hydrogen carbonate ion	
	1.1×10^{-3}	Hydrosulfuric acid	H,5	HS-	Hydrogen sulfide ion	
	6.3×10*	Dihydrogen phosphate ion	H ₂ PO ₄ -	HPO. *-	Hydrogen phosphate ion	
	6.2×10*	Hydrogen sulfite ion	HS-	50,*	Sulfite ion	
	2.9×10 ⁻⁴	Hypochlorous acid	HCIO	CIO-	Hypochlorite ion	
	6.2×10^{-10}	Hydrocyanic acid	HCN	ON-	Cyanide ion	
	5.8×10++	Ammonium ion	NH. +	NH ₃	Ammonia	
	5.8×10 **	Boric acid	н,во,	H ₂ BO ²	Dihydrogen borate	
	-4.7×10^{-11}	Hydrogen carbonate ion	HCO ₂ +	COal-	Carbonate ion	
	4.2×10^{-10}	Hydrogen phosphate ion	HPO, P	PO. F	Phosphate ion	
	$18\!\times\! 10^{-10}$	Dihydrogen borate ion	н,воз-	нво, н	Hydrogen borate ion	
	1.3×10^{-10}	Hydrogen sulfide ion	HS-	5.6	Sulfide Ion	
	1.6×10-**	Hydrogen borate ion	нво, -	BO ₃ *	Borate ion	
Weakest		water	H,O	OH-	Hydroxide	Stronges



Chemistry Acids And Bases Guide

M Tight

Chemistry Acids And Bases Guide:

Chemistry Acids And Bases Guide Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has be more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Chemistry Acids And Bases Guide**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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