

CHEMICAL QUANTITIES REVIEW QUESTIONS

- Calculate the following:
 - How many moles are present in 8.5 g of magnesium nitrate, $\text{Mg}(\text{NO}_3)_2$?
 - How many atoms are there in 8.5 g of $\text{Mg}(\text{NO}_3)_2$?
 - How many moles are there in 4.20×10^{21} molecules of methane?
 - What is the mass of 1.0×10^{23} molecules of water?
- Titanium reacts with oxygen to produce titanium oxide, the pigment that gives paint and sunscreen a white colour.
 - If 15.6 g of titanium reacts with oxygen to form 20.0 g of titanium oxide, calculate the mass of oxygen that reacted. What chemistry law are you using?
 - Calculate the percent composition of titanium oxide.
 - How much titanium oxide can be produced from 500 g of titanium and unlimited oxygen? What law are you using to solve this?
- What is the percent composition of:
 - sodium carbonate (Na_2CO_3)?
 - aluminum oxide (Al_2O_3)?
- What is the empirical formula of a compound that is:
 - 30.4% N, 69.6% O (by mass)
 - 43.6% P, 56.4% O (by mass)
- What is the molecular formula of a compound with molar mass 60.0 g and composed of 39.97% C, 6.73% H, and 53.30% O (by mass)?
- A certain compound has the following percent composition: C - 57.1 %, H - 4.8%, and O - 38.1 %. If the molar mass is 126 g, what is the molecular formula?
- When 25.0 g of a certain organic compound containing C, H and O is subjected to combustion analysis, 27.8 g of CO_2 and 19.9 g of H_2O are recovered. Determine the empirical formula of the compound.
- How many grams of lead (II) chloride will be produced when 6.7 g of lead (II) nitrate react with excess hydrochloric acid to form nitric acid and lead (II) chloride?
- A reaction involved in the production of iron from iron ore is:

$$\text{Fe}_2\text{O}_3 (\text{s}) + \text{CO} (\text{g}) \longrightarrow \text{Fe} (\text{s}) + \text{CO}_2 (\text{g})$$
 - How many kilograms of CO must react to produce 3.50 kg of Fe?
 - What mass of CO_2 would be produced?
- What is the maximum amount of iron (III) oxide than can be produced when 255 g of iron (II) sulfide are reacted with 71.4 g of oxygen gas in a closed container to produce iron (III) oxide and sulfur dioxide gas?
 - What mass of the excess reactant remains after the reaction?

Chemical Quantities Guided Practice Answers

M Mosston



Chemical Quantities Guided Practice Answers:

Chemical Quantities Guided Practice Answers Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chemical Quantities Guided Practice Answers**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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