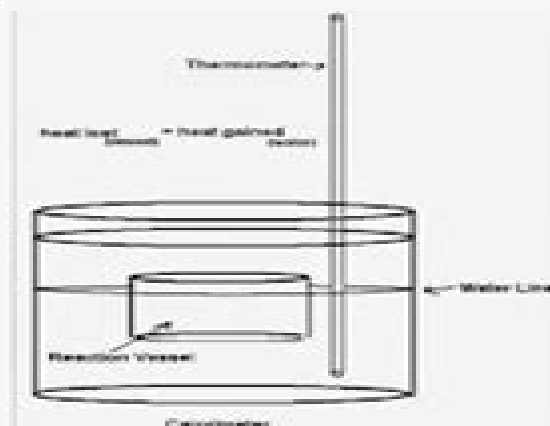


Model

Assume that a calorimeter is a closed system where all the energy released by an exothermic change is absorbed by the water in the calorimeter. If the mass of the water is known, the temperature change of the water can be used to determine the amount of heat energy released.



Equation for the calculation of heat is: $q = mC\Delta T$

q = heat released in Joules

m = mass of water in the calorimeter

C = specific heat capacity (**not $^{\circ}\text{C}$ which is a temperature**)

ΔT = [final temperature – initial temperature] (absolute value)

Key Questions:

1. What is the numerical value and units of the specific heat capacity of water?

Specific heat for water is $4.18 \frac{\text{J}}{\text{g}^{\circ}\text{C}}$ or $1 \frac{\text{cal}}{\text{g}^{\circ}\text{C}}$

2. What information does the specific heat capacity of water provide?

The specific heat value gives the amount of heat necessary to raise the temperature of one gram of water up 1°C .

3. How can the heat released into some mass of water be calculated from the specific heat capacity of water and the change in temperature of the water? Answer in words not with an equation.

Specific heat is how how much heat is needed to raise water 1°C , so if you how much the temperature changed, you can multiply the ΔT by specific heat to get the heat need to create temperature change.

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W. Zielenkiewicz, E. Margas



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Calorimetry of Non-Reacting Systems John P. McCullough, Donald W. Scott, 2013-10-22 Experimental Thermodynamics Volume 1 Calorimetry of Non Reacting Systems covers the heat capacity determinations for chemical substances in the solid liquid solution and vapor states at temperatures ranging from near the absolute zero to the highest at which calorimetry is feasible This book is divided into 14 chapters The first four chapters provide background information and general principles applicable to all types of calorimetry of non reacting systems The remaining 10 chapters deal with specific types of calorimetry Most of the types of calorimetry treated are developed over a considerable period and brought to a relatively sophisticated state For such calorimetry the approach adopted is to give detailed accounts of a few examples of apparatus and techniques representative of the best current practice in the field For the few types of calorimetry a general review of the field was considered more appropriate This book will prove useful to thermochemists engineers and experimentalists

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