

Advanced Placement Chemistry

1999 Multiple Choice Questions

Notes

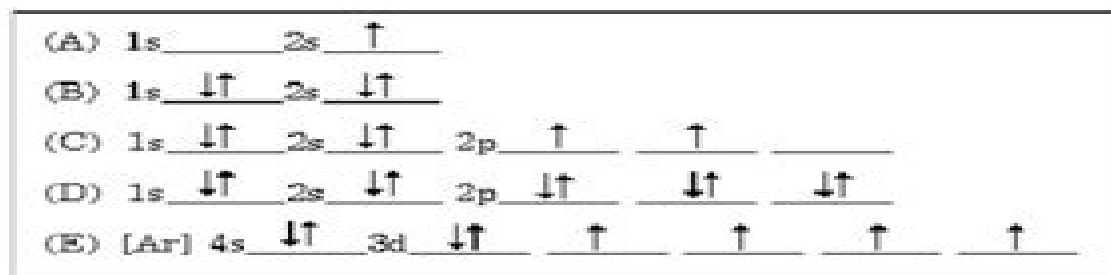
- Note: For all questions involving solutions and/or chemical equations, assume that the system is in pure water at room temperature unless otherwise noted.
- [Go to the answers](#)

Questions 1-4 refer to the following types of energy.

- (A) Activation energy
- (B) Free energy
- (C) Ionization energy
- (D) Kinetic energy
- (E) Lattice energy

1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion
2. The energy change that occurs in the conversion of an ionic solid to widely separated gaseous ions
3. The energy in a chemical or physical change that is available to do useful work
4. The energy required to form the transition state in a chemical reaction

Question 5-8 refer to atoms for which the occupied atomic orbitals shown below.



5. Represents an atom that is chemically unreactive

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