



20 Essential Chapters for

Contemporary Pharmacy Practice

THIRD EDITION

Judith E. Thompson

Professor, School of Pharmacy, University of Wisconsin

Contemporary Guide To Pharmacy Practice Thompson

Shelley Chambers Fox



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research methodology how to design successful experiments and ethics in science It also provides practical information on data collection and analysis keeping a laboratory notebook and writing laboratory reports With each section written by a distinguished researcher chapter 2 reviews common techniques for the characterization and analysis of colloidal structures including surface tension measurements viscosity and rheological measurements electrokinetic methods scattering and diffraction techniques and microscopy Chapters 3 5 provide 19 experiments each including the purpose of the experiment background information pre laboratory questions step by step procedures and post laboratory questions Chapter 3 contains experiments about colloids and surfaces such as sedimentation exploration of wetting phenomena foam stability and preparation of miniemulsions Chapter 4 covers various techniques for the preparation of nanoparticles including silver magnetic and silica nanoparticles Chapter 5 demonstrates daily life applications of colloid science describing the preparation of food colloids body wash and body cream

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enthalpy by simply summing up the changes for each step of the way until product is formed all steps have to proceed at the same temperature and the equations for the individual steps must balance out

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta H = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note

that the enthalpy change ΔH for products $\Delta H_{\text{products}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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