# Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO<sub>3</sub> (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

| Equation: | 2 KC10 <sub>3</sub> | → 2 KCI + | 3 02 |
|-----------|---------------------|-----------|------|
| Before:   | 1.2                 | •         | 0    |
| Change:   | -1.2                | +1.2      | +1.8 |
| After:    | 0                   | 1.2       | 1.8  |

1.2 moles KClO<sub>2</sub> needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

| Equations | Nig +  | 2 HCl →  | MgCl <sub>1</sub> + | HI a  |
|-----------|--------|----------|---------------------|-------|
| Before:   | 0.100  | 7        | 0                   | 0     |
| Change:   | -0.100 | -0.200 + | 0.100 +             | 0.100 |
| Afters    |        | 7        | 0.100               | 0.100 |

0.200 moles HCI needed

3. Ethane, C<sub>2</sub>H<sub>6</sub> reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

| Equation: | 2 C2H6 | + 7 O <sub>2</sub> | → 4 CO <sub>2</sub> | + 6.H <sub>2</sub> O |
|-----------|--------|--------------------|---------------------|----------------------|
| Before:   | 2.20   | 76.76              | 0                   |                      |
| Changes   | -2.20  | - 7.70             | + 4,40              | + 6.60               |
| After:    | 0      | 3.5                | 4.40                | 6.60                 |

O<sub>2</sub> needed: 7.70 mole x 32.0 g = 246 g O<sub>2</sub> required 1 mole

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:

Ni(NO<sub>3</sub>)<sub>2</sub> + 2 NaOH --> Ni(OH)<sub>2</sub> + 2 NaNO<sub>3</sub>

0.73 g x 1 mole = .0040 mole 183 g

| Equation: | Ni(NO <sub>3</sub> ): | 2 NaOH   | → Ni(OH)2 | + 2 NaNO3 |
|-----------|-----------------------|----------|-----------|-----------|
| Before:   | 0.0040                | XS       | 0         | 0         |
| Change:   | - 0.0040              | - 0.0080 | +0.0040   | +0.0080   |
| After:    | 0                     | NS       | 0.0040    | 0.0080    |

NaNO, produced: 0.0080 mole x 85.0 g = 0.68 g I mole

# **Chapter 5 Pearson Education Chemistry Answer Key**

Theodore L. Brown, H. Eugene LeMay
Jr., Bruce E. Bursten, Catherine
Murphy, Patrick Woodward, Steven
Langford, Dalius Sagatys, Adrian
George

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