

1. Here are three isotopes of an element:  $^{12}_6\text{C}$        $^{13}_6\text{C}$        $^{14}_6\text{C}$ 
  - a. The element is: **Carbon**
  - b. The number 6 refers to the: **Atomic Number**
  - c. The numbers 12, 13, and 14 refer to the **different mass numbers**
  - d. How many protons and neutrons are in the first isotope? **P: 6    N: 6**
  - e. How many protons and neutrons are in the second isotope? **P: 6    N: 7**
  - f. How many protons and neutrons are in the third isotope? **P: 6    N: 8**

2. Complete the following chart:

| Isotope name | atomic # | mass # | # of protons | # of neutrons | # of electrons |
|--------------|----------|--------|--------------|---------------|----------------|
| Potassium-37 | 19       | 37     | 19           | 18            | 19             |
| Oxygen-17    | 8        | 17     | 8            | 9             | 8              |
| uranium-235  | 92       | 235    | 92           | 143           | 92             |
| uranium-238  | 92       | 238    | 92           | 146           | 92             |
| boron-10     | 5        | 10     | 5            | 5             | 5              |
| boron-11     | 5        | 11     | 5            | 6             | 5              |

3. How do you calculate the mass number of an element?

**Mass # = protons + neutrons**

4. If an atom has 14 protons, 17 neutrons and 14 electrons, which expression identifies the element?
  - a. Silicon-14
  - b. Silicon-28
  - c. **Silicon-31**
  - d. Silicon-17
5. If an atom has 53 protons, 54 neutrons and 53 electrons, which expression identifies the element?
  - a. Iodine-53
  - b. Iodine-127
  - c. Iodine-54
  - d. **Iodine-107**

# Compass Learning Odyssey Answer Key Chemistry

**Rachel Turoscy**



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